

Chapter Test A

Chapter: Chemical Equations and Reactions

on the space provided below the letter of the item or phrase that best completes each statement or best answers each question.

1. The mass relationship in a chemical reaction is based on which of the following statements about atoms?
 - a. They combine to form molecules.
 - b. They are produced and destroyed.
 - c. They are in constant motion and are always in contact.
 - d. They do not change identity, mass, or chemical composition.
2. An equation described by a chemical equation involving x reactants and y products represents a balancing operation.
 - a. coefficient x is placed in front of products.
 - b. coefficient y is placed in front of reactants.
 - c. coefficient x is placed in front of reactants.
 - d. coefficient y is placed in front of products.
3. Which of the following relationships is described by the equation $aA + bB \rightarrow cC + dD$?
 - a. The products of a reaction are a molecules of A and b molecules of B , and the reactants are c molecules of C and d molecules of D .
 - b. All of the coefficients are the same.
 - c. The products of a reaction are a molecules of A and b molecules of B , and the reactants are c molecules of C and d molecules of D .
 - d. The reactants of a reaction are a molecules of A and b molecules of B , and the products are c molecules of C and d molecules of D .
4. There is a qualitative relationship between the number of molecules of a substance and the number of moles of that substance.
 - a. The number of molecules of a substance is equal to the number of moles of that substance.
 - b. The number of molecules of a substance is equal to the number of moles of that substance multiplied by the number of molecules in a mole.
 - c. The number of molecules of a substance is equal to the number of moles of that substance divided by the number of molecules in a mole.
 - d. All of the above.
5. An equation is a chemical equation if it is a chemical equation.
 - a. correct the coefficients.
 - b. correct the subscripts.
 - c. correct the number of molecules.
 - d. correct the number of atoms.

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