

Group 'A'

[11 × 1 = 11]

Multiple Choice questions

- The mass of anhydrous Na_2CO_3 required to prepare 250 ml of semi-normal solution is
a. 6.625g b. 13.265g c. 26.5g d. 53g
- The group reagent of Group II metal ions is
a. HCl b. $\text{HCl} + \text{H}_2\text{S}$
c. $\text{NH}_4\text{Cl} + \text{NH}_4\text{OH}$ d. $\text{NH}_4\text{Cl} + \text{NH}_4\text{OH} + \text{H}_2\text{S}$
- In a chemical reaction the product acts as catalyst. This phenomenon is
a. Enzyme Catalysis b. Homogenous Catalysis
c. Heterogenous Catalysis d. Autocatalysis
- In which of these case the reaction is always spontaneous?
a. $\Delta H = +ve, \Delta S = +ve$ b. $\Delta H = -ve, \Delta S = -ve$
c. $\Delta H = +ve, \Delta S = -ve$ d. $\Delta H = -ve, \Delta S = +ve$
- The configuration of d electron in Fe^{3+} in high spin complex
a. $d_{eg}^3 d_{eg}^2$ b. $d_{eg}^3 d_{eg}^0$ c. $d_{eg}^4 d_{eg}^1$ d. $d_{eg}^2 d_{eg}^3$
- Which of these is not an ore of copper?
a. Azurite b. Malachite c. Chalcopyrite d. Siderite
- 2-chloropropane undergoes dehydrohalogenation to give which of the following according to Saytzeff rule?
a. But-1-ene b. But-2-ene c. Propene d. Ethene
- The decreasing reactivity order of acid derivatives towards S_N reaction is:
a. Acid halides > Esters > Anhydrides > Amides
b. Anhydrides > Acid halides > Esters > Amides
c. Acid halides > Anhydrides > Esters > Amides
d. Anhydrides > Acid halides > Amides > Esters
- Mortar is the mixture of
a. Cement + water b. Cement + sand
c. Cement + lime d. Cement + sand + aggregate
- The process of enhancing physical and mechanical property of paper is called
a. Calendering b. Pulping c. Matting d. Glossing
- The radioisotope used for radiotherapy of cancer cells is
a. Fe-56 b. Co-60 c. I-131 d. P-30

Group 'B'

[8 × 5 = 40]

Short Answer questions

- Identify following as lewis acid or lewis base with reason i) BF_3 ii) NH_3
 - Calculate degree of ionization of 0.1M HCOOH solution and also calculate its pH (K_a $\text{HCOOH} = 1.8 \times 10^{-4}$)
- Or,
- Differentiate between enthalpy and internal energy.
 - Calculate heat of formation of glucose from the following data

Heat of combustion of glucose	- 2815 KJ/mole
Heat of formation of CO_2	- 395 KJ/mole
Heat of formation of H_2O	- 269 KJ/mole

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