IMFs are the forces that exist between the molecules, atoms, or ions that make up a particular substance

Stronger the IMFS, that harder it is to separate the particles

Fusion: solid → liquid and Vaporization: liquid → gas both require an input of energy

Condensation: gas→ liquid releases energy

 Enthalpy of vaporization ΔH_{vec} is the energy required to convert one mole of a liquid to vapor at a given temperature

As IMF † = ΔH.... †

11.2

 Vapor Pressure- the pressure of the vapor of a substance in contact with its liquid or solid phase in a sealed container

 When liquid is placed in a sealed container, some molecules escape into vapor phase (increasing vapor pressure) and some will recondense into liquid phase. When these become equal, they reach the dynamic equilibrium

 Equilibrium vapor pressure is the pressure of vapor of a substance at equilibrium in contact with its liquid or solid phase in a sealed container

† temperature = † VP

TIMEs=1 VP

Boiling Point of a liquid is when the Vapor pressure of the liquid is equal to the external pressure

 Normal boiling point is the temperature when the vapor pressure of a liquid is equal to 1 atm (760 mmHg)

†IMFs= †BP

$$\ln P = rac{-\Delta H_{
m vap}}{RT} + C \qquad \ln rac{P_2}{P_1} = rac{-\Delta H_{
m vap}}{R} \left(rac{1}{T_2} - rac{1}{T_1}
ight)$$

R= 8.314x10⁻³ kJ/K*mol, C=constant of integration

InP vs. 1/T and the slope is ΔH_m/R

11.3

 Cohesive forces: a force of attraction between molecules of a single substance and affect the physical properties of a liquid

· Surface tension is the measure of force required to "break" the surface of a liquid

 Results from different IMFs experienced by molecules at the surface of a liquid and those in the interior of the liquid

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 - The Rise of Digital Reading General Chemistry Semester 2 Study Guide
 - Advantages of eBooks Over Traditional Books
- 2. Identifying General Chemistry Semester 2 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - $\circ\,$ Features to Look for in an General Chemistry Semester 2 Study Guide
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from General Chemistry Semester 2 Study Guide
 - Personalized Recommendations
 - $\circ\,$ General Chemistry Semester 2 Study Guide User Reviews and Ratings
 - General Chemistry Semester 2 Study Guide and Bestseller Lists
- 5. Accessing General Chemistry Semester 2 Study Guide Free and Paid eBooks
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - o Carving Out Dedicated Reading Time
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 - Fact-Checking eBook Content of General Chemistry Semester 2 Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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