



A gaseous mixture contains 420.0 Torr $\text{H}_2(\text{g})$, 382.3 Torr $\text{N}_2(\text{g})$, and 67.1 Torr $\text{Ar}(\text{g})$. Calculate the mole fraction, x , of each of these gases.

$x_{\text{H}_2} =$

$x_{\text{N}_2} =$

$x_{\text{Ar}} =$

General Chemistry Mcquarrie Answers

A D Jenkins, John F Kennedy



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