

LAB REPORT

1) The two peaks of chlorine are 112 + 114.

molar mass of chlorobenzene = 112, 114

molar mass of benzene = 77

$$112 - 77 = 35$$

$$114 - 77 = 37$$

two isotopes
of Chlorine

The two peaks of bromine are 156 + 158.

$$156 - 77 = 79$$

$$158 - 77 = 81$$

two isotopes
of Bromine

~~The two peaks of bromine are~~

The natural abundances of chlorine - Cl = the lengths of 112 + 114 from the mass spectrum are 8.3 cm and 2.3 cm.

$$^{35}\text{Cl} = \frac{8.3}{8.3 + 2.3} = \frac{8.3}{10.6} \times 100\% = 78\%$$

$$^{37}\text{Cl} = \frac{2.3}{8.3 + 2.3} = \frac{2.3}{10.6} \times 100\% = 22\%$$

The natural abundances of bromine - Br = the lengths of 156 + 158 from the mass spectrum are 6.1 cm and 6.0 cm.

$$^{79}\text{Br} = \frac{6.1}{6.1 + 6.0} = \frac{6.1}{12.1} \times 100\% = 50\%$$

$$^{81}\text{Br} = \frac{6.0}{6.1 + 6.0} = \frac{6.0}{12.1} \times 100\% = 50\%$$

Nomenclature Report

1. The trend I noticed with observing the physical characteristics is when a compound was ionic and metal it formed a crystalline solid, but when the compound had a mixture of molecules or had larger molecules they formed amorphous solids.

Signature

Date

Witness/TA

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1. Understanding the eBook General Chemistry Laboratory Manual Chm 121
 - The Rise of Digital Reading General Chemistry Laboratory Manual Chm 121
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Laboratory Manual Chm 121
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Laboratory Manual Chm 121
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Laboratory Manual Chm 121
 - Personalized Recommendations
 - General Chemistry Laboratory Manual Chm 121 User Reviews and Ratings
 - General Chemistry Laboratory Manual Chm 121 and Bestseller Lists
5. Accessing General Chemistry Laboratory Manual Chm 121 Free and Paid eBooks
 - General Chemistry Laboratory Manual Chm 121 Public Domain eBooks
 - General Chemistry Laboratory Manual Chm 121 eBook Subscription Services
 - General Chemistry Laboratory Manual Chm 121 Budget-Friendly Options
6. Navigating General Chemistry Laboratory Manual Chm 121 eBook Formats

- ePub, PDF, MOBI, and More
- General Chemistry Laboratory Manual Chm 121 Compatibility with Devices
- General Chemistry Laboratory Manual Chm 121 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Laboratory Manual Chm 121
 - Highlighting and Note-Taking General Chemistry Laboratory Manual Chm 121
 - Interactive Elements General Chemistry Laboratory Manual Chm 121
- 8. Staying Engaged with General Chemistry Laboratory Manual Chm 121
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Laboratory Manual Chm 121
- 9. Balancing eBooks and Physical Books General Chemistry Laboratory Manual Chm 121
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Laboratory Manual Chm 121
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry Laboratory Manual Chm 121
 - Setting Reading Goals General Chemistry Laboratory Manual Chm 121
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry Laboratory Manual Chm 121
 - Fact-Checking eBook Content of General Chemistry Laboratory Manual Chm 121
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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