

Note that the end vowel of the Greek prefix is occasionally omitted from the spelling for clarity in pronunciation.

Hydrates

Prefixes are also used for the naming of **hydrates**, inorganic salts in which water molecules are a part of the crystalline structure of the solid. The waters of hydration (also called waters of crystallization) are bound to the ions of the salt and, in most cases, can be removed with the application of heat (see *Experiment 5*). For example, barium chloride is often purchased as barium chloride dihydrate: $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$. The formula of iron(III) chloride hexahydrate is $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$. The formula for Epsom salt (Figure 5.1) is $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$, magnesium sulfate heptahydrate.

Binary Acids

A binary acid is an *aqueous* solution of a compound formed by hydrogen and a more electronegative nonmetal. To name the acid, the prefix *hydro-* and suffix *-ic* are added to the root of the nonmetal name.

Formula	Name
HCl(aq)	<i>hydrochloric acid</i>
HBr(aq)	<i>hydrobromic acid</i>
$\text{H}_2\text{SO}_3\text{(aq)}$	<i>hydrosulfuric acid</i>

Writing Formulas from Names

In writing formulas for compounds, it is mandatory that the sum of the oxidation numbers of the elements equals zero (see *Dry Lab 2A, rule 8*). The name of the compound must dictate how the formula is to be written.

Example D2B.1 Write the formula for calcium nitride.

Calcium is in group 2A, and therefore has an oxidation number of $+2$: Ca^{2+} . The nitrogen atom is always -3 when combined with a metal: N^{3-} . Therefore, for the sum of the oxidation numbers to equal zero, there must be three Ca^{2+} (a total of $+6$) for every two N^{3-} (a total of -6); the formula is Ca_3N_2 .

DRY LAB PROCEDURE

Procedure Overview: Given the formula of the compound, the proper names for a large number of compounds are to be written. Given the name of the compound, the formulas for a large number of compounds also are to be written.

Your instructor will assign the exercises you are to complete. Answer them on a separate piece of paper. Be sure to indicate the date, your lab section, and your desk number on your *Report Sheet*. Use the rules that have been described.

1. Name the following binary salts of the representative elements.

- a. Na_3P d. CaC_2 g. Ca_3P_2 j. K_2S m. NH_4Br p. AlCl_3
 b. Na_2O e. CaI_2 h. KCN k. K_2Te n. $(\text{NH}_4)_2\text{S}$ q. Al_2O_3
 c. Na_3N f. CaH_2 i. KOH l. K_2O_2 o. NH_4CN r. AlN

2. Name the following salts according to the old *-ic*, *-ous* system and the Stock system.

- a. CrS g. HgCl_2 m. CoO
 b. Cr_2O_3 h. Hg_2Cl_2 n. $\text{CoBr}_2 \cdot 6\text{H}_2\text{O}$
 c. $\text{CrI}_3 \cdot 6\text{H}_2\text{O}$ i. HgO o. SnF_4
 d. CuCl j. Fe_2O_3 p. SnO_2
 e. CuI_2 k. FeS q. Cu_2O
 f. $\text{CuBr}_2 \cdot 4\text{H}_2\text{O}$ l. $\text{FeI}_3 \cdot 6\text{H}_2\text{O}$ r. $\text{Fe}(\text{OH})_3$

General Chemistry Lab Manual Answer Key

Scott C. Dulebohn



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Exploring General Chemistry in the Laboratory
Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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Introduction to General, Organic, and Biochemistry Morris Hein,1997 The authors recognize that both science and mathematics may be daunting subjects for many students taking this course With this in mind they have anticipated where students might stumble and have paced and organized this text to help them through Their goal is to make the material interesting and relevant so students understand the basic chemical principles related to their career The authors emphasize problem solving and provide a range of practice exercises As in previous editions the text first presents the basic concepts of general chemistry and then moves into organic and biochemistry In this edition the first two sections have been revised primarily to improve explanations and include new pedagogical features The biochemistry portion has been thoroughly updated to include coverage of many recent developments and emerging technologies in the field *Saunders Interactive*

General Chemistry John C. Kotz,William J. Vining,1998-09-21 An interactive presentation of general chemistry for college and university students While the Interactive Presentation makes up the majority of the material on these discs other items are ActivChemistry Software CAChe Visualizer for Education a tool for visualizing molecules and their properties Interactive periodic table database Database of Common Chemical Compounds Plotting tool Molecular weight and molarity calculators

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