

Chemistry 161 - EXAM REVIEW Handout (Lecture 13)

1. Aluminum (Al) is mined from aluminum oxide ( $\text{Al}_2\text{O}_3$ ), determine the mass of pure Al that could be extracted from a 467 g sample of  $\text{Al}_2\text{O}_3$ .

mass% of Al in  $\text{Al}_2\text{O}_3$ ?

$$\text{Al}_2\text{O}_3 = 101.96 \text{ g/mol}$$

$$26.98 \times 2 = 53.96 \text{ g/mol (mass of Al in } \text{Al}_2\text{O}_3)$$

$$\text{mass\%} = \frac{53.96 \text{ g}}{101.96 \text{ g}} \times 100 = 52.9\%$$



A sample of  $\text{Ag}_2\text{S}$  is found to contain  $2.64 \times 10^{24}$  atoms of Ag. What is the mass of the sample of  $\text{Ag}_2\text{S}$ ?

# General Chemistry 1 Final Exam

**RM Cervero**



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