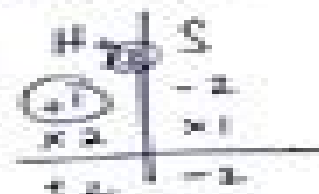
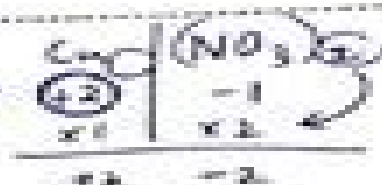


# Worksheet #3: Naming Inorganic Compounds

1. Give the name for each of the following ionic compounds:

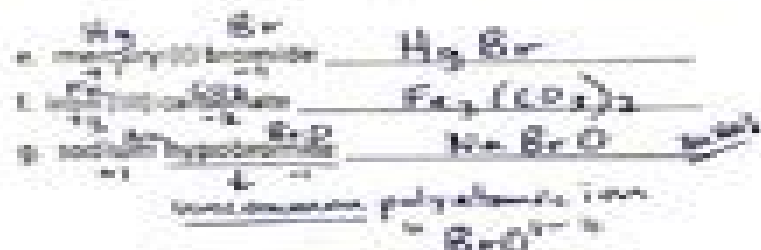
- $\text{AlF}_3$  Aluminum Fluoride
- $\text{Fe(OH)}_2$  Iron(II) Hydroxide
- $\text{Cu(NO}_3)_2$  Copper (II) Nitrate
- $\text{Ba(ClO}_3)_2$  Barium Perchlorate
- $\text{Li}_3\text{PO}_4$  Lithium Phosphate
- $\text{HgS}$  Mercury (I) Sulfide
- $\text{Ca(CH}_3\text{COO)}_2$  Calcium Acetate
- $\text{Cr}_2(\text{CO}_3)_3$  Chromium (III) Carbonate
- $\text{K}_2\text{Cr}_2\text{O}_7$  Potassium Dichromate
- $\text{(NH}_4)_2\text{SO}_4$  Ammonium Sulfate



all of the underlined compounds are polyatomic ions  $\rightarrow$  Memorize Them

2. Write the chemical formula for each of the following compounds:

- copper(II) oxide  $\text{Cu}_2\text{O}$
- potassium peroxide  $\text{K}_2\text{O}_2$
- aluminum hydroxide  $\text{Al(OH)}_3$
- iron(III) bromide  $\text{FeBr}_3$



3. Give the name or chemical formula, as appropriate, for each of the following acids:

- nitric acid  $\text{HNO}_3$
- hydrobromic acid  $\text{HBr}$
- phosphoric acid  $\text{H}_3\text{PO}_4$

- hypochlorous acid  $\text{HClO}$
- iodic acid  $\text{HIO}_3$
- sulfurous acid  $\text{H}_2\text{SO}_3$

Band off of Iodate is  $\text{IO}_3^-$

4. Give the name or chemical formula, as appropriate, for each of the following molecular substances:

- iodine heptafluoride  $\text{IF}_7$
- iodine pentafluoride  $\text{IF}_5$
- Xenon Trioxide  $\text{XeO}_3$

- dinitrogen tetroxide  $\text{N}_2\text{O}_4$
- hydrogen cyanide  $\text{HCN}$
- tetraphosphorus hexasulfide  $\text{P}_4\text{S}_6$

There are all binary ionic compounds (+/-)

0 - None  
1 - One  
2 - Two  
3 - Three  
4 - Four  
5 - Five  
6 - Six  
7 - Seven  
8 - Eight  
9 - Nine

# Gen Chem Unit 3 Packet Answers

**Annelies Wilder-Smith**



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