



When the substances in the equation above are at equilibrium at pressure P and temperature T, the equilibrium can be shifted to favor the products by

- (A) increasing the pressure by means of a moving piston at constant T
- (B) increasing the pressure by adding an inert gas such as nitrogen
- (C) decreasing the temperature
- (D) allowing some gases to escape at constant P and T
- (E) adding a catalyst

2. In which of the following systems would the number of moles of the substances present at equilibrium NOT be shifted by a change in the volume of the system at constant temperature?

- (A) $\text{CO(g)} + \text{NO(g)} \rightleftharpoons \text{CO}_2\text{(g)} + 1/2 \text{ N}_2\text{(g)}$
- (B) $\text{N}_2\text{(g)} + 3 \text{ H}_2\text{(g)} \rightleftharpoons 2 \text{ NH}_3\text{(g)}$
- (C) $\text{N}_2\text{(g)} + 2 \text{ O}_2\text{(g)} \rightleftharpoons 2 \text{ NO}_2\text{(g)}$
- (D) $\text{N}_2\text{O}_4\text{(g)} \rightleftharpoons 2 \text{ NO}_2\text{(g)}$
- (E) $\text{NO(g)} + \text{O}_3\text{(g)} \rightleftharpoons \text{NO}_2\text{(g)} + \text{O}_2\text{(g)}$

3. At constant temperature, a change in volume will NOT affect the moles of substances present in which of the following?

- (A) $\text{H}_2\text{(g)} + \text{I}_2\text{(g)} \rightleftharpoons 2\text{HI(g)}$
- (B) $\text{CO(g)} + \text{Cl}_2\text{(g)} \rightleftharpoons \text{COCl}_2\text{(g)}$
- (C) $\text{PCl}_5\text{(g)} \rightleftharpoons \text{PCl}_3\text{(g)} + \text{Cl}_2\text{(g)}$
- (D) $\text{N}_2\text{(g)} + 3 \text{ H}_2\text{(g)} \rightleftharpoons 2 \text{ NH}_3\text{(g)}$
- (E) $\text{CH}_4\text{(g)} + \text{CO}_2\text{(g)} \rightleftharpoons 2 \text{ CO(g)} + 2 \text{ H}_2\text{(g)}$

4. $\text{PCl}_5\text{(g)} + \text{Cl}_2\text{(g)} \rightleftharpoons \text{PCl}_3\text{(g)} + \text{energy}$
 Some PCl_5 and Cl_2 are mixed in a container at 200°C and the system reaches equilibrium according to the equation above. Which of the following causes an increase in the number of moles of PCl_5 present at equilibrium?

- I. Decreasing the volume of the container
- II. Raising the temperature
- III. Adding a mole of He gas at constant volume

- (A) I only
- (B) II only
- (C) I and III only
- (D) II and III only
- (E) I, II, and III



When 0.40 mole of SO_2 and 0.60 mole of O_2 are placed in an evacuated 1.00-liter flask, the reaction represented above occurs. After the reactants and the product reach equilibrium and the initial temperature is restored, the flask is found to contain 0.30 mole of SO_3 . Based on these results, the expression for the equilibrium constant, K_c , of the reaction is

- (A) $(0.30)^2 / [(0.45)(0.10)^2]$
- (B) $(0.30)^2 / [(0.60)(0.40)^2]$
- (C) $(2 \times 0.30) / [(0.45)(2 \times 0.10)]$
- (D) $(0.30) / [(0.45)(0.10)]$
- (E) $(0.30) / [(0.60)(0.40)]$

6. How many moles of NaF must be dissolved in 1.00 liter of a saturated solution of PbF_2 at 25°C to reduce the $[\text{Pb}^{2+}]$ to 1×10^{-8} molar?
 (K_{sp} of PbF_2 at $25^\circ\text{C} = 4.0 \times 10^{-8}$)
- (A) 0.020 mole
 - (B) 0.040 mole
 - (C) 0.10 mole
 - (D) 0.20 mole
 - (E) 0.40 mole

7. The solubility of CuI is 2×10^{-8} molar. What is the solubility product constant, K_{sp} , for CuI ?
 (A) 1.4×10^{-17}
 (B) 2×10^{-16}
 (C) 4×10^{-12}
 (D) 2×10^{-12}
 (E) 8×10^{-18}

8. What is the molar solubility in water of Ag_2CrO_4 ?
 (The K_{sp} for Ag_2CrO_4 is 8×10^{-12} .)
 (A) $8 \times 10^{-12} \text{ M}$
 (B) $2 \times 10^{-12} \text{ M}$
 (C) $(4 \times 10^{-12} \text{ M})^{1/2}$
 (D) $(4 \times 10^{-12} \text{ M})^{1/3}$
 (E) $(2 \times 10^{-12} \text{ M})^{1/3}$

9. A student wishes to reduce the zinc ion concentration in a saturated zinc iodate solution to $1 \times 10^{-6} \text{ M}$. How many moles of solid KIO_3 must be added to 1.00 L of solution? ($K_{sp} \text{ Zn(IO}_3)_2 = 4 \times 10^{-6}$ at 25°C)

- (A) 1 mol
- (B) 0.5 mol
- (C) 0.0001 mol
- (D) 4 mol
- (E) 2 mol

Equilibrium Test 2 Multiple Choice Answers

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