Equilibrium Practice Test

101210112220011	Name: ANSWI	ERS		60 /60
witiple Choice cutify the letter of the choice that i	house committees when see	and the same of th	the execution (2)	O mortal
	And a second-second second			
1. Equilibrium hoppens when				010010000000000000000000000000000000000
a, opposing forces belance b. opposing forces are unbal		c, the forward and reverse reactions occur at different rates d, the forward and reverse reactions are unbalanced		
an appropriate tracks are contain		a. a	E TES SE STAR STAR	The Control of the Co
2. Which is a macroscopic pro-	pentity 7			
a, colone		c. concentration		
b. temperature		d. pH		
3. A system reaches equilibrius	m. Which most be?	Completes with he seem		
a. it is open. c. it is either open or closed				
b, it is closed		d. it is open and has	backware.	
 Ice cobes are meiting in a gl system? 	ass of water with a	lid on it. Why does	this situation (not represent an equilibrius
a. It is not closed system.				
b. The equilibrium cannot be		ne opposite dispetion	is the second second	
The temperature is not cos The rates of opposing ch		arc .		
or a sec a sec of approximation		-		
5. Which of these is NOT an er	comple of beteroger			
a. H ₂ O(1)eedH ₂ O(g)		c. N₂O₂(g)_{***}2NO₃(
b. $C(x) + CO_3(g) e + CO_3(g)$		d. C ₂ H ₂ OH(t)+++C ₃ H	GOHigs	
6. In the equilibrium constan	it expression below,	what do the expen-	ents represent	•
(C) (D)				
$K_{eq} = \frac{[C]^n [D]^n}{(A)^n B(^n)}$				
a. the concentrations of reac	tumbs and products	c. the rates of the fo-	eward and rever	nse reactions
b. the equilibrium constant		d. the stolchlometr	tic coefficients.	from the balanced equation
7. Which is tree of the nemeri-	cal value of the equ	differiwan constant, b	C7	
a. It represents the relative	concentration of p	moducts compared:	witth reactionts.	
b. It represents the ratio of r				
c. It is measured in the SI us	sit, mole.			
d. Both a and c are true.				
8. What does it mean if $K_{**} < 1$	L at equilibrium?			
a. The concentration of re-		on the concentration	on of products.	
b. The concentration of prod	lucts in greater thus t	the concentration of r	neac Sanda.	
c. The rate of the forward re-				
st. The rate of the reverse rea	actions in greater than	. The rate of the forwa	and epactions.	
9. When the equilibrium const the initial concentr		he approximate con	scentration of	reactant at equilibrium is
as greater than		c. the same as		
b. less than		d. exactly 100 times		
10. If Q _{et} > K _{ef} then				
a, the system is at equalibrium	_	at the section	and proposition the first	reactants is greater than t
b. the ratio of products to re-				d product formation
11. Use the following informati				
The following reaction is an	example of homoge	meous equilibrium.		
$H_0(g) + I_0(g) \leftrightarrow 2HI(g)$				
The state of the s		The second secon	The second secon	Committee of the Commit
If the pressure is hadved, b	y doubling the volu	me and the temper c, not be affected	ature remains	constant, the equilibrium w

12. Which condition is NOT a condition for an equilibrium system?

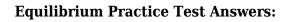
a. equal rates for opposing changes

b. a closed system

d. constant temperature

Equilibrium Practice Test Answers

Julia Schneider



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