

**Multiple Choice**

Identify the letter of the choice that best completes the statement or answers the question. (20 marks)

- Equilibrium happens when \_\_\_\_\_.  
**a. opposing forces balance**  
 b. opposing forces are unbalanced  
 c. the forward and reverse reactions occur at different rates  
 d. the forward and reverse reactions are unbalanced
- Which is a macroscopic property?  
**a. colour**  
 b. temperature  
 c. concentration  
 d. pH
- A system reaches equilibrium. Which must be true of the system?  
**b. It is closed**  
 a. it is open  
 c. it is either open or closed  
 d. it is open and has pressure
- Ice cubes are melting in a glass of water with a lid on it. Why does this situation not represent an equilibrium system?  
**d. The rates of opposing changes are not equal.**  
 a. It is not closed system.  
 b. The equilibrium cannot be approached from the opposite direction.  
 c. The temperature is not constant.
- Which of these is NOT an example of heterogeneous equilibrium?  
**c.  $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$**   
 a.  $\text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_2\text{O}(\text{g})$   
 b.  $\text{C}(\text{s}) + \text{CO}_2(\text{g}) \rightleftharpoons \text{CO}(\text{g})$   
 d.  $\text{C}_2\text{H}_5\text{OH}(\text{l}) \rightleftharpoons \text{C}_2\text{H}_5\text{OH}(\text{g})$
- In the equilibrium constant expression below, what do the exponents represent?  

$$K_{eq} = \frac{[\text{C}]^b [\text{D}]^d}{[\text{A}]^a [\text{B}]^c}$$
  
**d. the stoichiometric coefficients from the balanced equation**  
 a. the concentrations of reactants and products  
 b. the equilibrium constant  
 c. the rates of the forward and reverse reactions
- Which is true of the numerical value of the equilibrium constant,  $K_{eq}$ ?  
**a. It represents the relative concentration of products compared with reactants.**  
 b. It represents the ratio of reactants to products.  
 c. It is measured in the SI unit, mole.  
 d. Both a and c are true.
- What does it mean if  $K_{eq} < 1$ , at equilibrium?  
**a. The concentration of reactants is greater than the concentration of products.**  
 b. The concentration of products is greater than the concentration of reactants.  
 c. The rate of the forward reaction is greater than the rate of the reverse reaction.  
 d. The rate of the reverse reaction is greater than the rate of the forward reaction.
- When the equilibrium constant is very small, the approximate concentration of reactant at equilibrium is \_\_\_\_\_ the initial concentration.  
**c. the same as**  
 a. greater than  
 b. less than  
 d. exactly 100 times
- If  $Q_{eq} > K_{eq}$ , then \_\_\_\_\_.  
**c. the ratio of products to reactants is greater than  $K_{eq}$**   
 a. the system is at equilibrium  
 b. the ratio of products to reactants is less than  $K_{eq}$   
 d. the reaction shifts toward product formation
- Use the following information to answer the next question.  
 The following reaction is an example of homogeneous equilibrium:  

$$\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$$
  
**c. not be affected**  
 If the pressure is halved, by doubling the volume and the temperature remains constant, the equilibrium will  
 a. shift forward.  
 b. shift backward  
 d. will not reach equilibrium
- Which condition is NOT a condition for an equilibrium system?  
**c. reactants in the same phase**  
 a. equal rates for opposing changes  
 b. a closed system  
 d. constant temperature

# Equilibrium Practice Test Answers

**Julia Schneider**



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