

Chapter

7

THERMOCHEMISTRY

MULTIPLE CHOICE QUESTIONS

- Spontaneous reactions are: (a) Reversible (b) Irreversible (c) Non irreversible (d) None of these **(LHR 05)**
- The standard heat changes occur at: (a) 25°C and 2 atm (b) 298 K and 1 atm (c) 25°C and 1 mm hg (d) 273 K and 1 atm **(BWP 08, GRW 06)**
- $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ the change in enthalpy for reaction is called: (a) Heat of reaction (b) heat of formation (c) Heat of neutralization (d) Heat of combustion **(LHR 06)**
- The net change in energy in a chemical reaction is same whether it takes place directly or indirectly. It is called: (a) Henry's law (b) Charlie's law (c) Hess's law (d) Graham's law **(RWP 11, GRW 07, BWP 10)**
- The enthalpy change when one mole of substance is completely burnt in excess of oxygen is called: (a) Enthalpy of atomization (b) Enthalpy of neutralization (c) Enthalpy of Combustion (d) Enthalpy of formation **(LHR 07)**
- Enthalpy change for the reaction: $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ is called enthalpy of: (a) Formation (b) Combustion (c) Neutralization (d) Atomization **(LHR 08)**
- Standard enthalpies are measured at: (a) 273 K (b) 298K (c) 373 K (d) All of these **(FSD 11, LHR 09)**
- The exothermic process is: (a) Evaporation (b) Sublimation (c) Respiration (d) Boiling **(DGK 10, GRW, LHR 11)**
- A state function which describes together the internal energy and the product of pressure and volume is called: (a) Enthalpy (b) Internal energy (c) Work (d) Free energy **(GRW 11)**

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Jin-Ying Zhang



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Table of Contents Multiple Choice Chemistry Questions On Thermochemistry

1. Understanding the eBook Multiple Choice Chemistry Questions On Thermochemistry
 - The Rise of Digital Reading Multiple Choice Chemistry Questions On Thermochemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Chemistry Questions On Thermochemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Chemistry Questions On Thermochemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Chemistry Questions On Thermochemistry

- Personalized Recommendations
 - Multiple Choice Chemistry Questions On Thermochemistry User Reviews and Ratings
 - Multiple Choice Chemistry Questions On Thermochemistry and Bestseller Lists
5. Accessing Multiple Choice Chemistry Questions On Thermochemistry Free and Paid eBooks
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 - Multiple Choice Chemistry Questions On Thermochemistry Budget-Friendly Options
 6. Navigating Multiple Choice Chemistry Questions On Thermochemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Multiple Choice Chemistry Questions On Thermochemistry Compatibility with Devices
 - Multiple Choice Chemistry Questions On Thermochemistry Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Choice Chemistry Questions On Thermochemistry
 - Highlighting and Note-Taking Multiple Choice Chemistry Questions On Thermochemistry
 - Interactive Elements Multiple Choice Chemistry Questions On Thermochemistry
 8. Staying Engaged with Multiple Choice Chemistry Questions On Thermochemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Chemistry Questions On Thermochemistry
 9. Balancing eBooks and Physical Books Multiple Choice Chemistry Questions On Thermochemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Chemistry Questions On Thermochemistry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Multiple Choice Chemistry Questions On Thermochemistry
 - Setting Reading Goals Multiple Choice Chemistry Questions On Thermochemistry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Multiple Choice Chemistry Questions On Thermochemistry

- Fact-Checking eBook Content of Multiple Choice Chemistry Questions On Thermochemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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