- A) Determine the percent composition as
 - B) Determine the percent composition of all elements in magnesium nitrate:

 - A) A compound is composed of 85.7 % Carbon and 14.3 % Hydrogen. Determine the Empirical Formula.

compound is composed of 85.7 % Carbon and 14.3 % Hydrograph (1.08/7.0% = 1)

$$85.79 \times \frac{1 \text{ mol}}{12.015} = 7.08 \text{ mol}$$
 $14.16/7.08 = 2 = C1/2$
 $14.39 \times \frac{1 \text{ mol}}{1.01} = 14.16 \text{ mol}$
 $14.16/7.08 = 2 = C1/2$

B) The empirical formula for a compound is SO₃. If the Molecular Mass of a compound is 239.9 grams,

Molar Mess of
$$503 = 803/m$$
 = 5309 = 5309 = 5309

C) An unknown compound contains 54.68% Carbon, 9.20 % Hydrogen, and 36.14% Oxygen. If the

D) Determine the empirical formula for the following compound. It contains 4.372 grams of lithium, 13.44

E) is it possible for a compound to have the same empirical formula and molecular formula? Explain and give an example.

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Pamela Mills, Amina El-Ashmawy

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