

## 17: Molecular Geometry - VSEPR

### Key Structure Terms

- Valence Shell:** Electrons in the outermost shell that are involved in bonding.
- Lewis Structure:** A 2D representation of a molecule and its bonds.
- Lone Pair:** Pair of electrons not being shared in a bond.
- Bonding Pair:** Pair of electrons used to make a bond. Both atoms sharing the electrons can "count" them in their valence shell.
- Single bond:** One pair of shared electrons.
- Double bond:** Two pairs of shared electrons. Shorter and stronger than a single bond.
- Triple bond:** Three pairs of shared electrons. Shorter and stronger than a double bond.
- Valence Shell Electron Pair Repulsion Theory (VSEPR):** Bonds and lone pairs are electrons, all electrons are negative, negative things repel other negative things, therefore all bonds and lone pairs arrange themselves in 3D as far away from each other as possible.
- Electron Geometry:** 3D structure of a molecule determined by counting the electron regions around a central atom (bonds and lone pairs).
- Electron Region:** Each bond (single, double or triple) and lone pair count as "1" electron region.
- Molecular Geometry:** 3D structure determined by the atoms bonded to the central atom.
- Ligand:** Atoms bonded to the central atom.

### VSEPR Theory and Geometry

**Electron geometry** is determined by looking at the number of electron regions around the central atom.

**Molecular geometry** is determined by looking at the number of atoms bonded to the central atom (ligand) and the number of lone pairs around the central atom.

**Electronic Geometry Mnemonic:** Linear(2) – Trigonal Planar(3) – Tetrahedral(4) – Trigonal Bipyramidal(5) – Octahedral(6) = "Long DIP To TRIBE Overseas."

**Examples:**



Electron geometry: Tetrahedron  
Molecular geometry: Tetrahedron



Electron geometry: Tetrahedron  
Molecular geometry: Trigonal pyramidal

### Molecular Geometries

A = central atom; X = ligands; E = lone pairs

Electron Regions	Molecular Formula	Name	Shape
2	$AX_2$ (BeCl <sub>2</sub> , CO <sub>2</sub> )	Linear	
3	$AX_3$ (BF <sub>3</sub> , CO <sub>3</sub> <sup>2-</sup> , NO <sub>3</sub> <sup>-</sup> , SO <sub>3</sub> )	Trigonal Planar	
	$AX_2E$ (NO <sub>2</sub> <sup>+</sup> , SO <sub>2</sub> , O <sub>3</sub> )	Bent	
4	$AX_4$ (CH <sub>4</sub> , NH <sub>4</sub> <sup>+</sup> , PO <sub>4</sub> <sup>3-</sup> , SO <sub>4</sub> <sup>2-</sup> , ClO <sub>4</sub> <sup>-</sup> )	Tetrahedron	
	$AX_3E$ (NH <sub>3</sub> , H <sub>3</sub> O <sup>+</sup> , PCl <sub>3</sub> , SO <sub>3</sub> <sup>2-</sup> )	Trigonal pyramidal	
	$AX_2E_2$ (H <sub>2</sub> O, ClO <sub>2</sub> , OF <sub>2</sub> , SCl <sub>2</sub> )	Bent	
5	$AX_5$ (PCl <sub>5</sub> )	Trigonal bipyramidal	
	$AX_4E$ (SF <sub>4</sub> , SCl <sub>5</sub> )	See-saw	
	$AX_3E_2$ (ClF <sub>3</sub> , ICl <sub>3</sub> )	T-shaped	
	$AX_2E_3$ (XeF <sub>2</sub> , I <sub>3</sub> <sup>-</sup> )	Linear	
6	$AX_6$ (SF <sub>6</sub> , PCl <sub>6</sub> <sup>-</sup> )	Octahedron	
	$AX_5E$ (BrF <sub>5</sub> , IF <sub>5</sub> )	Square pyramidal	
	$AX_4E_2$ (XeF <sub>4</sub> )	Square planar	

**How to Use This Cheat Sheet:** These are the keys related to this topic. Try to read through it carefully twice then write it out on a blank sheet of paper. Review it again before the exams.

# Molecular Structure And Bonding Cheat Sheet

**Kaplan Test Prep**

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

## **Molecular Structure And Bonding Cheat Sheet:**

**Organic Chemistry I For Dummies** Arthur Winter, 2016-05-13 Organic Chemistry I For Dummies 2nd Edition 9781119293378 was previously published as Organic Chemistry I For Dummies 2nd Edition 9781118828076 While this version features a new Dummies cover and design the content is the same as the prior release and should not be considered a new or updated product The easy way to take the confusion out of organic chemistry Organic chemistry has a long standing reputation as a difficult course Organic Chemistry I For Dummies takes a simple approach to the topic allowing you to grasp concepts at your own pace This fun easy to understand guide explains the basic principles of organic chemistry in simple terms providing insight into the language of organic chemists the major classes of compounds and top trouble spots You ll also get the nuts and bolts of tackling organic chemistry problems from knowing where to start to spotting sneaky tricks that professors like to incorporate Refreshed example equations New explanations and practical examples that reflect today s teaching methods Fully worked out organic chemistry problems Baffled by benzines Confused by carboxylic acids Here s the help you need in plain English

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## **Table of Contents Molecular Structure And Bonding Cheat Sheet**

1. Understanding the eBook Molecular Structure And Bonding Cheat Sheet
  - The Rise of Digital Reading Molecular Structure And Bonding Cheat Sheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Molecular Structure And Bonding Cheat Sheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molecular Structure And Bonding Cheat Sheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Molecular Structure And Bonding Cheat Sheet
  - Personalized Recommendations
  - Molecular Structure And Bonding Cheat Sheet User Reviews and Ratings
  - Molecular Structure And Bonding Cheat Sheet and Bestseller Lists
5. Accessing Molecular Structure And Bonding Cheat Sheet Free and Paid eBooks
  - Molecular Structure And Bonding Cheat Sheet Public Domain eBooks
  - Molecular Structure And Bonding Cheat Sheet eBook Subscription Services

- Molecular Structure And Bonding Cheat Sheet Budget-Friendly Options
- 6. Navigating Molecular Structure And Bonding Cheat Sheet eBook Formats
  - ePub, PDF, MOBI, and More
  - Molecular Structure And Bonding Cheat Sheet Compatibility with Devices
  - Molecular Structure And Bonding Cheat Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Molecular Structure And Bonding Cheat Sheet
  - Highlighting and Note-Taking Molecular Structure And Bonding Cheat Sheet
  - Interactive Elements Molecular Structure And Bonding Cheat Sheet
- 8. Staying Engaged with Molecular Structure And Bonding Cheat Sheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Molecular Structure And Bonding Cheat Sheet
- 9. Balancing eBooks and Physical Books Molecular Structure And Bonding Cheat Sheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Molecular Structure And Bonding Cheat Sheet
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Molecular Structure And Bonding Cheat Sheet
  - Setting Reading Goals Molecular Structure And Bonding Cheat Sheet
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molecular Structure And Bonding Cheat Sheet
  - Fact-Checking eBook Content of Molecular Structure And Bonding Cheat Sheet
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web here is the form otherwise let us wish you all the best for the coming youth occasion in the church and we want to pray to our heavenly father to grant you the grace to help you to the end we are praying for and the blessings of the lord be abundantly showered upon your life through jesus christ

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web i greet you in the name of our lord jesus allow me to thank our lord jesus and our father god for making it possible for us to congregate today here for this very special occasion in our church when we are commemorating this 10 years of service in the church

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web i greet you in the name of the lord jesus good morning what a joy to have you attending this occasion in our church that is organized by the youths we have been praying for this days and want to thank god for enabling us to be here today

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web i greet you in the name of our lord jesus on behalf of the church and the department of youth i want to take this opportunity to thank our god for allowing us to see this day it is a day as a department we have been praying for and all the youths of this church are happy that we are here to fellowship together

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web jul 29 2023 express your gratitude towards the guests for attending by using a phrase such as it s so wonderful to see you all here on this sunny day 2 for an event with close friends and family more informal language might be appropriate include a few jokes and keep your speech more lighthearted

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