

Name _____ Date: _____

The Mole Webquest

Use the following website to answer this set of questions:

<http://antoine.frostburg.edu/chem/scenesc/101/moles/faq/why-use-moles.shtml>

- 1) A mole of anything is how many? (give the number):

$$6.02 \times 10^{23}$$

- 2) Why is it that different amounts of things can still equal one mole? (think about the weight of a dozen elephants vs a dozen eggs)

The number of grams in a mole is different from substance to substance.

- 3) Why do we want to use the concept of moles?

You often want to know how many molecules you have in a sample of a substance.

With moles, you count the number of molecules in the sample by weighing it.

- 4) Once we know the number of moles we can convert to the number of:

moles or grams and vice versa.

- 5) How many grams of water are in one mole of water?

$$1 \text{ mole water} = 18 \text{ g water}$$

- 6) How many molecules of water are in one mole of water?

$$6.02 \times 10^{23}$$

Use the following website to answer the next set of questions:

<https://chemfiesta.wordpress.com/2015/02/06/mole-calculations-3/>

- 7) Define molar mass.

The mass of a given substance (chemical element or chemical compound is in g) divided by its amount of substance in (mol).

- 8) What is the mass of one mole of oxygen?

16.0 grams per mole

- 8) Find the molar mass of NH_4NO_3 . (Show your work)

Ammonium nitrate has two nitrogen atoms ($2 \times 14.0 \text{ g/mol}$), four hydrogen atoms ($4 \times 1.0 \text{ g/mol}$), and three oxygen atoms ($3 \times 16.0 \text{ g/mol}$). This makes the molar mass 80.0 g/mol .

Mole Webquest Key

CL Gary

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'CL Gary' text.

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Moles University of Florida. Agricultural Experiment Station, J.R. Watson, 1945

Adventures of Mole, 1996

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