

Name \_\_\_\_\_ Date: \_\_\_\_\_

### The Mole Webquest

Use the following website to answer this set of questions:

<http://antoine.frostburg.edu/chem/scenesc/101/moles/faq/why-use-moles.shtml>

- 1) A mole of anything is how many? (give the number):

$$6.02 \times 10^{23}$$

- 2) Why is it that different amounts of things can still equal one mole? (think about the weight of a dozen elephants vs a dozen eggs)

**The number of grams in a mole is different from substance to substance.**

- 3) Why do we want to use the concept of moles?

**You often want to know how many molecules you have in a sample of a substance.**

**With moles, you count the number of molecules in the sample by weighing it.**

- 4) Once we know the number of moles we can convert to the number of:

**moles or grams** and vice versa.

- 5) How many grams of water are in one mole of water?

$$1 \text{ mole water} = 18 \text{ g water}$$

- 6) How many molecules of water are in one mole of water?

$$6.02 \times 10^{23}$$

Use the following website to answer the next set of questions:

<https://chemfiesta.wordpress.com/2015/02/06/mole-calculations-3/>

- 7) Define molar mass.

**The mass of a given substance (chemical element or chemical compound is in g) divided by its amount of substance in (mol).**

- 8) What is the mass of one mole of oxygen?

**16.0 grams per mole**

- 8) Find the molar mass of  $\text{NH}_4\text{NO}_3$ . (Show your work)

**Ammonium nitrate has two nitrogen atoms ( $2 \times 14.0 \text{ g/mol}$ ), four hydrogen atoms ( $4 \times 1.0 \text{ g/mol}$ ), and three oxygen atoms ( $3 \times 16.0 \text{ g/mol}$ ). This makes the molar mass  $80.0 \text{ g/mol}$ .**

# Mole Webquest Answer Key

**Ying Liu**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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