

Model 2 – Proposed Calculations for Mass of NH_3 to Mass of N_2

Toby's Method

$$\frac{x \text{ grams}}{50.0 \text{ g}} = \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3} \rightarrow x = \underline{15} \text{ g } \text{N}_2$$

Rachel's Method

$$50.0 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} = \underline{2.94} \text{ moles } \text{NH}_3$$

$$\frac{x \text{ mole } \text{N}_2}{\underline{1.5} \text{ mole } \text{NH}_3} = \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3} \rightarrow x = \underline{0.11} \text{ moles } \text{N}_2$$

$$\underline{0.11} \text{ mole } \text{N}_2 \times \frac{28.0 \text{ g } \text{N}_2}{1 \text{ mole } \text{N}_2} = \underline{2.43} \text{ g } \text{N}_2$$

Jerry's Method

$$50.0 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} \times \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3} \times \frac{28.0 \text{ g } \text{N}_2}{1 \text{ mole } \text{N}_2} = \underline{2.43} \text{ g } \text{N}_2$$

- Model 2 shows three proposed calculations to solve the problem in Question 10. Complete the calculations in Model 2 by filling in the underlined values.
- Which method does not use the mole ratio in an appropriate manner? Explain.
Toby's method does not correctly use the mole ratio that is required.
- Two of the methods in Model 2 give the same answer. Show that they are mathematically equivalent methods.

$$\begin{aligned} &\text{Jerry's method} \\ &\left\{ \begin{aligned} &50 \times \frac{1}{17} = \frac{1}{2} \times \frac{28}{1} = 2.43 \text{ g } \text{N}_2 \\ &\quad \quad \quad \downarrow \\ &\quad \quad \quad \frac{1}{2} \times \frac{1}{2} \\ &\quad \quad \quad \downarrow \\ &0.11 \times \frac{28}{1} = 2.43 \text{ g } \text{N}_2 \end{aligned} \right. \end{aligned}$$

- Use either Rachel or Jerry's method from Model 2 to calculate the mass of hydrogen needed to make 50.0 g of ammonia. $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

$$50.0 \text{ g} \left(\frac{1 \text{ mole}}{17.0 \text{ g}} \right) \left(\frac{3 \text{ moles}}{2 \text{ moles}} \right) \left(\frac{1 \text{ g}}{2.016 \text{ g}} \right) = 2.43 \text{ g } \text{H}_2 \text{ required}$$

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Scientific and Technical Aerospace Reports ,1994 Chemical Process Design and Simulation: Aspen Plus and Aspen Hysys Applications Juma Haydary,2019-01-23 A comprehensive and example oriented text for the study of chemical process design and simulation Chemical Process Design and Simulation is an accessible guide that offers information on the most important principles of chemical engineering design and includes illustrative examples of their application that uses simulation software A comprehensive and practical resource the text uses both Aspen Plus and Aspen Hysys simulation software The author describes the basic methodologies for computer aided design and offers a description of the basic steps of process simulation in Aspen Plus and Aspen Hysys The text reviews the design and simulation of individual simple unit operations that includes a mathematical model of each unit operation such as reactors separators and heat exchangers The author also explores the design of new plants and simulation of existing plants where conventional chemicals and material mixtures with measurable compositions are used In addition to aid in comprehension solutions to examples of real problems are included The final section covers plant design and simulation of processes using nonconventional components This important resource Includes information on the application of both the Aspen Plus and Aspen Hysys software that enables a comparison of the two software systems Combines the basic theoretical principles of chemical process and design with real world examples Covers both processes with conventional organic chemicals and processes with more complex materials such as solids oil blends polymers and electrolytes Presents examples that are solved using a new version of Aspen software ASPEN One 9 Written for students and academics in the field of process design Chemical Process Design and Simulation is a practical and accessible guide to the chemical process design and simulation using proven software **NASA Technical Note** ,1967 Chemistry James N. Spencer,George M. Bodner,Lyman H. Rickard,2010-12-28 Chemistry Structure and Dynamics 5th Edition emphasises deep understanding rather than comprehensive coverage along with a focus on the development of inquiry and reasoning skills While most mainstream General Chemistry texts offer a breadth of content coverage the Spencer author team in contrast focuses on depth and student preparation for future studies The fifth edition is revised in keeping with our commitment to the chemical education community and specifically the POGIL Process Oriented Guided Inquiry Learning Project This text reflects two core principles first that the concepts that are covered are fundamental building blocks for understanding chemistry and second that the concepts should be perceived by the students as being directly applicable to their interests and careers The authors further provide this core coverage using 1 of 3 models data driven chemical theories and student understanding which allows for a more concrete foundation on which students build conceptual understanding **Official Gazette of the United States Patent and Trademark Office** ,2000 *Power Systems and Power Plant Control* 1989 U. Ahn,2014-06-05 The control of power systems and power plants is a subject of growing interest which continues to sustain a high level of research development and application in many diverse

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Official Gazette of the United States Patent Office United States. Patent Office,1968

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