MOLARITY BY DILUTION

Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid = M x V then, $M_1 \times V_1 = M_2 \times V_2$. Solve the following problems.

1. How much concentrated 18 M sulfuric acid is needed to prepare 250 mL of a 6.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(6.0 M)(.250 L)}{18 M}$$

$$\boxed{83 mL}$$

2. How much concentrated 12 M hydrochloric acid is needed to prepare 100. mL of a 2.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(2.0 M)(100 mL)}{12 M} =$$

$$\boxed{17 mL}$$

3. To what volume should 25 mL of 15 M nitric acid be diluted to prepare a 3.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(15 M)(25 mL)}{3.0 M}$$

$$\boxed{125 mL}$$

4. To how much water should 50. mL of 12 M hydrochloric acid be added to produce a 4.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(12 M)(50 mL)}{4.0 M} = 150 mL - 50 mL$$

$$\boxed{100 mL\ H_2 O}$$

5. To how much water should 100. mL of 18 M sulfuric acid be added to prepare a 1.5 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(18 M)(100 mL)}{1.5 M} = 1200 mL - 100 mL = 1100 mL$$

$$\boxed{1.1 L}$$

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