

Stoichiometry Worksheet

Name _____

Date _____

Solve the following problems. Show ALL work and include the units in both the setup and the final answer.

1. How many grams of NaCl are needed to prepare 1.00 liter of a 0.100 molar solution?
2. What mass of MgSO_4 is needed to prepare 0.50 liter of a 0.200 molar solution?
3. The gaseous form of H_2SO_4 is SO_3 . What mass of sulfur is required?
4. What mass of iron(III) sulfate is needed to make 500. mL of a 0.25 M solution?
5. How many grams of iron sulfate are needed to prepare 200 mL of a 0.100 M solution?
6. How many grams of sodium are needed to prepare 2.00 L of a 0.200 M solution of NaOH ?
7. What mass of potassium sulfate is needed to prepare 0.500 L of a 0.750 M solution of this salt.
8. What mass of CaCl_2 is needed to prepare 0.200 L of a 0.750 M solution?
9. If 1.50 grams of Ca(OH)_2 are needed to prepare 1.00 mL of solution, what will be the molarity of the solution?
10. What is the molarity of a solution of 100. mL of solution contains 0.50 g of Na_2SO_4 ?
11. What is the molarity of a solution that contains 0.50 g of sodium, and is 750 mL of solution?
12. What is the molarity of a solution that contains 0.50 g of NaCl in 250 mL of solution?
13. What is the molarity of three (3) 500 mL beakers of NaOH in 1.00 liter of solution?
14. Find the concentration in molarity of a solution that contains 0.75 g acetone ($\text{C}_3\text{H}_6\text{O}$) in a solution of 0.750 L.
15. What volume of solution would be prepared if 0.010 g CaCl_2 would be used to make a 0.00010 solution?

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