

The sum of the pressure of the H_2 inside the tube (mm Hg), plus the equivalent in mm Hg of the height of the water remaining in the tube above the surface of the water in the beaker, plus the vapor pressure (mm Hg) of the water at the temperature of the water in the tube (assume it is the same as T in beaker water) must equal the atmospheric pressure pressing down on the surface of the water outside the tube (Figure 1c).

$$P_{\text{atm}} (\text{torr}) = P_{H_2} (\text{torr}) + P_{H_2O} (\text{vapor}) + P (h \text{ column } H_2O, \text{ torr})$$

The vapor pressure of water at the temperature measured can be found in the back of your laboratory procedure manual, or in the "Handbook of Chemistry and Physics" "CRC". The pressure of the column of water in torr (mm Hg) will be equal to the mm H_2O divided by 13.59, a correction obtained by multiplying the height of the water column by the ratio of the density of water to that of mercury.

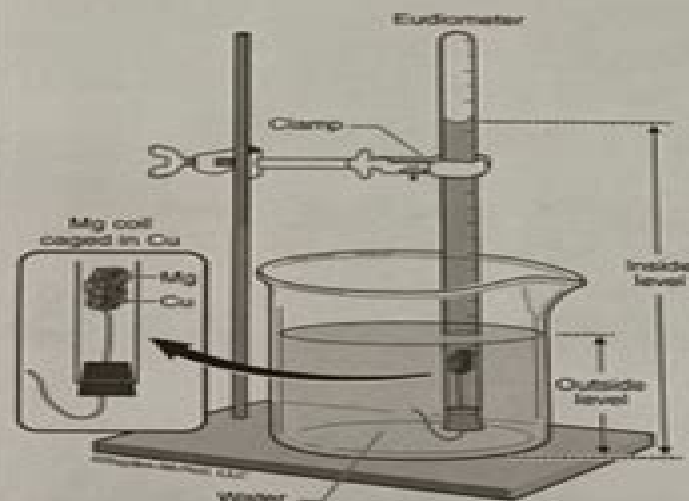


Figure 1

The appropriate volume of concentrated (12 M HCl) is added to the buret, and water is then layered on top of it until the buret is completely filled, as in Figure 1a. Inversion of the buret in a beaker partly filled with water starts the reaction (1b), which continues until all of the metal is gone and the buret is nearly full of hydrogen gas, as shown in 1c.

Molar Volume Of Hydrogen Gas Lab Answers

**United States. National Aeronautics
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