

Heat Energy Changes in Chemical Reactions

1 - Which two words are used to describe heat energy changes in chemical reactions?

- a) Thermal and Exothermal
- b) Endoscopic and Endothermal
- c) Exothermic and Endothermic
- d) Exoskeleton and Endoskeleton

2 - In which type of reaction is heat energy taken in from the surroundings?

- a) Exothermic
- b) Endothermic
- c) Nuclear
- d) Dehydration

3 - In which type of reaction is heat energy given out to the surroundings?

- a) Exothermic
- b) Endothermic
- c) Nuclear
- d) Dehydration

4 - Which piece of lab equipment would you use to measure the temperature of a chemical reaction?

- a) Gas syringe
- b) Indicator paper
- c) Measuring cylinder
- d) Thermometer

5 - In a reaction between magnesium and hydrochloric acid the following temperature change was observed:

Starting Temperature = 23°C

End Temperature = 29°C

What does this tell us about the reaction?

- a) The reaction is endothermic
- b) The reaction is exothermic
- c) The rate of reaction is very fast
- d) This is a neutralisation reaction

6 - In chemical reactions, bonds between atoms are broken and new bonds are formed. In both processes energy is involved. Which of these best describes bond breaking and bond making in terms of energy?

- a) Bond breaking and bond making both require energy to be input
- b) Bond making releases energy and bond breaking requires energy to be put in
- c) Bond making requires energy to be put in and bond breaking releases energy
- d) Bond breaking and bond making both releases energy

Modern Chemistry Reaction Energy Multiple Choice Answers

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