

- Chapter 5 Assessment Problems pp. 146-147 #33, 34, 39, 44, 45, 47, 49, 50, 52, 53, 56, 57, 59, 61, 63, 64, 65-72, 76, 78, 79, 80, 81(a-d,f), 82.
- 33. Explain how the gaseous neon atoms in a neon sign emit light.
Electricity passing through the tube excites electrons in the neon atoms to higher energy levels. When the excited electrons relax, they emit light as they lose energy.
- 34. What is a photon?
A particle of electromagnetic radiation having a mass of zero and carrying a quantum of energy.
- 39. What is the difference between an atom's ground state and an excited state?
The ground state is the lowest energy state, while any energy state higher than that is excited.
- 44. What is an atomic orbital?
A three-dimensional region around the nucleus describing an electron's probably location.
- 45. What is the probability that an electron will be found within an atomic orbital? 90%
- 47. How many energy sublevels are contained in each of the hydrogen atoms first three energy levels?
Level 1: 1; level 2: 2; level 3: 3
- 49. Which of the following atomic orbital designations are incorrect?
a. 7f b. 3f c. 2d d. 6p b & c are incorrect—the orbitals do not exist
- 50. What do the designations s, p, d, and f specify with respect to the atom's orbitals?
The shapes of the orbitals.
- 52. What is the maximum number of electrons an orbital may contain? 2
- 53. Why is it impossible to know precisely the velocity and position of an electron at the same time?
The photon required to measure an electron's velocity or position changes both the position and velocity of the electron (like a bowling ball thrown at a ping pong ball).
- 56. How is an orbital's principal quantum number related to the atom's major energy levels?
They are the same.
- 57. Explain the meaning of the Aufbau principle as it applies to atoms with many electrons.
The Aufbau principle describes the sequence in which an atom's orbitals fill with electrons.
- 59. Why must the two arrows within a single block of an orbital diagram be written in opposite (up and down) directions?
They must have opposite spins.
- 61. What are valence electrons? How many of a magnesium atom's 12 electrons are valence electrons?
The electrons in an atom's outermost orbitals. Magnesium has 2 electrons.
- 63. Describe the difference between a quantum and a photon.
A quantum is the minimum amount of energy that can be lost or gained by an atom, while a photon is a particle of light that carries a quantum of energy.

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