No. of Electron Pairs	Electron Pair Geometry (Bond Angle)	No. of Pendant Atoms	Molecular Geometry		mple Image
2	linear (180°)	2	linear	BeH <sub>2</sub>	
3	trigonal planar (120°)	3	trigonal planar	CO <sub>3</sub> 2-	-2-
		2	bent	NO <sub>2</sub> -	-
4	tetrahedral (109.5°)	4	tetrahedral	CH <sub>4</sub>	-32
	(107 deg)	3	trigonal pyramidal	NH <sub>3</sub>	30
	(104.5 deg)	2	bent	H <sub>2</sub> O	-
5	trigonal bipyramidal (90°, 120°)	5	trigonal bipyramidal	PCl <sub>5</sub>	-1-
		4	see-saw	SF <sub>4</sub>	
		3	T-shaped	BrF <sub>3</sub>	
		2	linear	ICl2-	• • •
6	octahedral (90°)	6	octahedral	SF <sub>6</sub>	4
		5	square pyramidal	BrF <sub>5</sub>	-
		4	square planar	ICl <sub>4</sub> -	> <

# **Modern Chemistry Ch 8 Review In Answers**

James F. Hall

# Modern Chemistry Ch 8 Review In Answers:

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\*\*Lab Experiments Modern Chemistry\*\* Harold Clark Metcalfe, 1982

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