Chapter 4 &5: Modern Model of the Atom and Bonding

Leader: Katie
Course: Chem 163
Instructor: Burnett/Kingston
Date: September 28, 2009

Supplemental Instruction lowa State University

Review

- 1. What is the noble gas configuration of the following atoms and ions?
 - a) Ar [Ne] 352 3pf or [Ar]
 - b) Nan+DAO 25 20 5 Blo]
 - e) C^2-[He] 26-204
 - d) 0"2-[He] 25" 2p" [NE]
 - e) Car2+ Lue > 35 50 60 CAr3

New Material

- 1. Define the following terms and answer associated questions.
 - a) Heisenberg uncertainty principle: Principle that states that the position and momentum of a subatanic particle connot be known simultaneously.
 - b) orbital: 310 volume in space, cloud like distribution of electrons.

-What is its maximum capacity? ⊇ €-

c) covalent bond: Chemical bond where the electrons one shored.

-What does it represent? The covalent-band represents on evergy minimum,
-Is energy absorbed or released when a covalent bond is formed? Evergy is released.

d) bond length:

-What causes bond length? + the attraction of electrons to the nuclei controps and fine attraction of electrons to the nuclei controps and fine attraction of electrons to the nuclei controps and fine attraction of electrons are the nuclei control of the nuclei con

2. How many orbitals are there in any s sublevel? Any p sublevel? Any d sublevel?

S=X Ichital - 2cp = 3orbitals - 6ea = Sorbitals - 10e-

3. Draw the orbitals for the 2p^5. How many electrons are there in each orbital?

8 35 3



- 4. Write the following atoms in Lewis Dot notation.
 - a) Nb) H
- 63 NJ-

O Xe

- c) Xe d) Fe
- @ H.

⊕ Fe

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TEAS 6 Exambusters

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- 1. Understanding the eBook Modern Biology Study Guide Answer Key Biochemistry
 - The Rise of Digital Reading Modern Biology Study Guide Answer Key Biochemistry
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Modern Biology Study Guide Answer Key Biochemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Biology Study Guide Answer Key Biochemistry
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Modern Biology Study Guide Answer Key Biochemistry
 - Personalized Recommendations
 - Modern Biology Study Guide Answer Key Biochemistry User Reviews and Ratings
 - Modern Biology Study Guide Answer Key Biochemistry and Bestseller Lists
- 5. Accessing Modern Biology Study Guide Answer Key Biochemistry Free and Paid eBooks
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- Modern Biology Study Guide Answer Key Biochemistry Budget-Friendly Options
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 - Modern Biology Study Guide Answer Key Biochemistry Enhanced eBook Features
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 - Highlighting and Note-Taking Modern Biology Study Guide Answer Key Biochemistry
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- 8. Staying Engaged with Modern Biology Study Guide Answer Key Biochemistry
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Biology Study Guide Answer Key Biochemistry
- 9. Balancing eBooks and Physical Books Modern Biology Study Guide Answer Key Biochemistry
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Modern Biology Study Guide Answer Key Biochemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Biology Study Guide Answer Key Biochemistry
 - Setting Reading Goals Modern Biology Study Guide Answer Key Biochemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Biology Study Guide Answer Key Biochemistry
 - Fact-Checking eBook Content of Modern Biology Study Guide Answer Key Biochemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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