

## Chapter 9 Stoichiometry Chapter Review Answers

1. Mole Ratios
- Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
  - They are formed from the coefficients found in a balanced chemical equation.
2. Mole Ratios as conversion factors
- $\frac{2 \text{ mol Cu}}{1 \text{ mol O}_2} \cdot \frac{2 \text{ mol CuO}}{2 \text{ mol Cu}} = \frac{2 \text{ mol CuO}}{1 \text{ mol O}_2}$ . Remember the ratios can be reduced as well!!
  - $\frac{1 \text{ mol O}_2}{2 \text{ mol Cu}} \cdot \frac{2 \text{ mol CuO}}{2 \text{ mol Cu}} = \frac{2 \text{ mol CuO}}{1 \text{ mol O}_2}$
  - $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{2 \text{ mol MgF}_2}{1 \text{ mol Mg}}$
  - $\frac{2 \text{ mol HF}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} = \frac{2 \text{ mol HF}}{1 \text{ mol MgF}_2}$
3.  $\frac{160.0 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} \cdot \frac{1 \text{ mol Na}_2\text{CO}_3}{74.10 \text{ g Ca(OH)}_2} = \frac{160.0 \text{ g Na}_2\text{CO}_3}{74.10 \text{ g Ca(OH)}_2}$
4. Molar Mass
- Definition: the mass of one mole of a substance
  - Molar mass is used as a conversion factor in stoichiometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
5. Hydrogen + Oxygen  $\rightarrow$  Water
- 5.0 mol H<sub>2</sub>
  - 2.5 mol O<sub>2</sub>
6. Ethane combustion reaction
- 15.8 mol O<sub>2</sub>
  - 9.00 mol CO<sub>2</sub> & 13.5 mol H<sub>2</sub>O
7. 575 g Na & 886 g Cl<sub>2</sub>
8. Iron production reaction
- 75.1 mol CO
  - 50.1 mol Fe & 75.1 mol CO<sub>2</sub>
9. 87,390 g CO or  $8.739 \times 10^4$  g CO & 12,610 g H<sub>2</sub> or  $1.261 \times 10^4$  g H<sub>2</sub>
10. Lightning reaction of N<sub>2</sub> & O<sub>2</sub>
- 1,100 g NO<sub>x</sub> or  $1.10 \times 10^3$  g NO<sub>x</sub>
  720. g NO or  $7.20 \times 10^2$  g NO
11. CO<sub>2</sub> removal in spacecraft
- 42.04 mol NaOH
  - 21.02 mmol Na<sub>2</sub>CO<sub>3</sub> & 21.02 mol H<sub>2</sub>O
12. Double replacement AgNO<sub>3</sub> & NaBr
- 463 g NaBr
  - 845 g AgBr
13. Soda-acid fire extinguisher
- 3.058 mol NaHCO<sub>3</sub>
  - 3.058 mol CO<sub>2</sub>, 1.529 mol Na<sub>2</sub>SO<sub>4</sub>, 3.058 mol H<sub>2</sub>O
14. Sulfuric acid-sodium hydroxide neutralization reaction
- H<sub>2</sub>SO<sub>4</sub> + 2 NaOH  $\rightarrow$  Na<sub>2</sub>SO<sub>4</sub> + 2 H<sub>2</sub>O
  - 37 g H<sub>2</sub>SO<sub>4</sub>
  - 53 g Na<sub>2</sub>SO<sub>4</sub> & 14 g H<sub>2</sub>O
15. Copper & silver nitrate reaction
- 0.0104 mol Cu(NO<sub>3</sub>)<sub>2</sub>
  - 0.0104 mol Cu & 0.0209 mmol AgNO<sub>3</sub>
16. Aspirin production
- 13.5 kg C<sub>8</sub>H<sub>8</sub>O<sub>3</sub>
  - 7.66 kg C<sub>8</sub>H<sub>8</sub>O<sub>3</sub>
  - 4.29 L H<sub>2</sub>C<sub>6</sub>H<sub>5</sub>O<sub>2</sub>
17. Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
18. The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant limits the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

# Modern Chemistry Stoichiometry chapter Review

## Answers

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