

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds.

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20 \\ + & & \\ \hline & & 63.62 \rightarrow 64 \end{array}$$

Modern Chemistry Answer Review

JL Elias



Modern Chemistry Answer Review:

Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **Reviews of Modern Quantum Chemistry** Kali Das Sen, 2002 This important book collects together state-of-the-art reviews of diverse topics covering almost all the major areas of modern quantum chemistry. The current focus in the discipline of chemistry synthesis, structure, reactivity, and dynamics is mainly on control. A variety of essential computational tools at the disposal of chemists have emerged from recent studies in quantum chemistry. The acceptance and application of these tools in the interfacial disciplines of the life and physical sciences continue to grow. The new era of modern quantum chemistry throws up promising potentialities for further research. Reviews of Modern Quantum Chemistry is a joint endeavor in which renowned scientists from leading universities and research laboratories spanning 22 countries present 59 in-depth reviews. Along with a personal introduction written by Professor Walter Kohn, Nobel laureate in Chemistry, 1998, the articles celebrate the scientific contributions of Professor Robert G. Parr on the occasion of his 80th birthday. List of Contributors: W. Kohn, M. Levy, R. Pariser, B. R. Judd, E. Lo, B. N. Plakhutin, A. Savin, P. Politzer, P. Lane, J. S. Murray, A. J. Thakkar, S. R. Gadre, R. F. Nalewajski, K. Jug, M. Randić, G. Del Re, U. Kaldor, E. Eliav, A. Landau, M. Ehara, M. Ishida, K. Toyota, H. Nakatsuji, G. Maroulis, A. M. Mebel, S. Mahapatra, R. Carbó, C. Dorca, U. Nagy, I. A. Howard, N. H. March, S. O. Collier, B. Liu, R. G. Pearson, N. Watanabe, S. Ten-no, S. Iwata, Y. Udagawa, E. Valderrama, X. Fradera, I. Silanes, J. M. Ugalde, R. J. Boyd, E. V. Ludea, V. V. Karasiev, L. Massa, T. Tsuneda, K. Hirao, J. M. Tao, J. P. Perdew, O. V. Gritsenko, M. Gruning, E. J. Baerends, F. Aparicio, J. Garza, A. Cedillo, M. Galvin, R. Vargas, E. Engel, A. Hack, R. N. Schmid, R. M. Dreizler, J. Poater, M. Sola, M. Duran, J. Robles, X. Fradera, P. K. Chattaraj, A. Poddar, B. Maiti, A. Cedillo, S. Gutierrez, O. Coliva, P. Jaque, A. Toro, C. Labb, R. H. Chermette, P. Boulet, S. Portmann, P. Fuentealba, R. Contreras, P. Geerlings, F. De Proft, R. Balawender, D. P. Chong, A. Vela, G. Merino, F. Kootstra, P. L. de Boeij, R. van Leeuwen, J. G. Snijders, N. T. Maitra, K. Burke, H. Appel, E. K. U. Gross, M. K. Harbola, H. F. Hameka, C. A. Daul, I. Ciofini, A. Bencini, S. K. Ghosh, A. Tachibana, J. M. Cabrera, O. Co-Trujillo, F. Tenorio, O. Mayorga, M. Cases, V. Kumar, Y. Kawazoe, A. M. Koster, P. Calaminici, Z. Gmez, U. Reveles, J. A. Alonso, L. M. Molina, M. J. Lopez, F. Dugue, A. Maanes, C. A. Fahlstrom, J. A. Nichols, D. A. Dixon, P. A. Derosa, A. G. Zacarias, J. M. Seminario, D. G. Kanhere, A. Vichare, S. A. Blundell, Z. CoY, Lu, H. CoY, Liu, M. Elstner, W. CoT, Yang, J. Muoz, X. Fradera, M. Orozco, F. J. Luque, P. Tarakeshwar, H. M. Lee, K. S. Kim, M. Valiev, E. J. Bylaska, A. Gramada, J. H. Weare, J. Brickmann, M. Keil, T. E. Exner, M. Hoffmann. Probability Distributions and Valence Shells in Atoms: A. Savin. Information Theoretical Approaches to Quantum Chemistry: S. R. Gadre. Quantum Chemical Justification for Clar's Valence Structures: M. Randić. Functional Expansion Approach in Density Functional Theory: S. B. Liu. Norm-conserving Pseudopotentials for the Exact Exchange Functional: E. Engel et al. Volume II: Chemical Reactivity and Dynamics within a Density-based Quantum Mechanical Framework: P. K. Chattaraj et al. Fukui Functions and Local Softness: H. Chermette et al. The Nuclear Fukui Function: P. Geerlings et al. Causality in Time-Dependent Density Functional Theory: M. K. Harbola. Theoretical Studies of Molecular Magnetism: H. F. Hameka. Melting in Finite Sized Systems: D. G. Kanhere et al. Density

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