

Name: \_\_\_\_\_  
 Period: \_\_\_\_\_

## Metals vs. Non-Metals; Dot Diagrams; Ions

Ch. 19.1

### Metals versus Non-Metals

The dividing line is the red or bold black line from between Boron and Aluminum down and to the right. Everything to the left is a metal; to the right, non-metal. One exception (don't ya just hate that?!) is hydrogen – a non-metal.

Sodium (Na) is a metal.

Oxygen is a non-metal.

Metal or Non-metal?	
Potassium: _____	Fluorine: _____
Bromine: _____	Hydrogen: _____
Beryllium: _____	Silver: _____
Helium: _____	Nitrogen: _____

*Metals are on the left side. Non-metals on the right.  
 Metals tend to lose electrons. Non-metals gain them tight.*

1 H									2 He
3 Li	4 Be			5 B	6 C	7 N	8 O	9 F	10 Ne
11 Na	12 Mg			13 Al	14 Si	15 P	16 S	17 Cl	18 Ar
19 K	20 Ca			31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr

Non-metal (exception)

Dividing line

Non-Metals

Metals

Easy to remember by which side iron (Fe) is on—the left side!

On either side of the divide are the metalloids or semi-metals—they have characteristics of both B, Si, Ge, As.

### Dot Diagrams

Dot Diagrams (sometimes known as Lewis dot diagrams) are a depiction of an atom's valence electrons. They are a powerful tool in helping you understand, see, and even predict molecular bonding.

Put the correct number of valence electrons around the chemical symbol in this order in pairs:



Neon has 8 valence electrons (in the right-hand column).  
 Neon has no unoccupied spaces.  
 It is full.

Notice that the electrons are in pairs, not a circle.

Why would the electrons spread out around the atom instead of bunching up?

Draw the Dot Diagrams for the following elements:

Beryllium

Sodium

Helium

Carbon

Boron



Lithium, with 1 valence electron (first column).

Carbon has 4 valence electrons

Oxygen has 6 valence electrons.  
 How many more before it's full?  
 \_\_\_\_\_

# Metals Vs Nonmetals Dot Diagrams Ions Worksheet

## Answer Key

**Sir Tarun Rupani**



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## **Table of Contents Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key**

1. Understanding the eBook Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
  - The Rise of Digital Reading Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
  - Personalized Recommendations
  - Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key User Reviews and Ratings
  - Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key and Bestseller Lists
5. Accessing Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key Free and Paid eBooks
  - Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key Public Domain eBooks
  - Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key eBook Subscription Services
  - Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key Budget-Friendly Options
6. Navigating Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key eBook Formats

- ePub, PDF, MOBI, and More
  - Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key Compatibility with Devices
  - Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
  - Highlighting and Note-Taking Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
  - Interactive Elements Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
8. Staying Engaged with Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
9. Balancing eBooks and Physical Books Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
- Setting Reading Goals Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
- Fact-Checking eBook Content of Metals Vs Nonmetals Dot Diagrams Ions Worksheet Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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