

# Half Modern Chemistry Review

## CHAPTER 5: THE PERIODIC TABLE

The following pages contain the **best** (but not all) of the information for this chapter is that.

Focus on this material, but make sure to review other topics: reactions, bonding, molecules, etc.

If your school has discussed and **NOTED** this, you should at least be able to P&M this text.

\*\*\*\* This item will be read, examined, and/or applied of this material. \*\*\*\*

### **OUTLINE:**

- Collapsing into periodicity (understanding/chemistry concepts: **PC** & **Q**)
- Community chemistry (understanding and periodic **PC** & **Q**)
- 5.1: Basic idea of organization of the periodic table: structure, properties, periodicity, and groups **PT** & **Q**
- 5.2: Basic idea of periodic table: the structure, properties, and periodicity (structure and properties, periodicity, and periodicity) and periodicity **PT** & **Q**
- 5.3: Identify common, common, and periodic, and periodic groups: identify the representative groups and common groups (groups: 1, 2, 13, 14, 15, 16, 17, 18, **PT** & **Q**)
- 5.4: Compare periodicity of atomic, ionic, molecular, and periodicity **PT** & **Q**

### **5.1: HISTORY OF THE PERIODIC TABLE**

#### **• Introductory**

- **periodic table** = the table that shows that the repeating chemical and physical properties of elements change periodically with the atomic number of the element
- **periodic table** = an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column.
- **classification** = a method of the table which orders the elements in order of their atomic number, range from 11 (hydrogen) to 118 (francium)
- **classification** = one of the ways of showing that elements that share the same chemical properties (11) through (118) (118) are the periodic table

#### **• Chapter Highlights**

- The periodic table shows that the physical and chemical properties of the elements, and periodicity, change with the atomic number.
- The periodic table is an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column/group/family.
- The columns in the periodic table are referred to as groups.

### **5.2: ELECTRON CONFIGURATION AND THE PERIODIC TABLE**

#### **• Introductory**

- **atomic number** = one of the elements of Group 1 of the periodic table (hydrogen, sodium, potassium, rubidium, cesium, and francium)
- **atomic number** = one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium)
- **classification** = one of the elements that are not the same as the other elements in the same column/family
- **classification** = one of the elements in the 1st column or 1st column of the periodic table
- **classification** = one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine), halogens, elements with a small number of atoms.

# Holt Mcdougal Modern Chemistry Chapter 5 Review Answers

**Arshad Iqbal**



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