

REPORT SHEET

Determination of the K_a for a Weak Acid

Name _____

Partner's Name _____

Date _____

Section _____

Instructor _____

(Except where otherwise noted, 3 pts. for all blanks.)

Efficiency of pH meter _____ (3 pts.)

NaOH Molarity 0.001 M (3 pts.)

Since the titration with 1-mL increments is done to give a rough estimate of the equivalence point, you need report on only the results from the more precise increment titration.

Acetic AcidLiterature value for K_a of Acetic Acid _____ 1.8×10^{-5} (3 pts.)Literature value for pK_a of Acetic Acid _____4.75

Volume of base to reach equivalence point _____

20.4 mL

pH at equivalence point _____

9. pH = 7

Volume of base halfway to equivalence point _____

10.2

pH at halfway to equivalence point _____

5.6Experimental pK_a of Acetic Acid _____

(3 pts.)

Percent error of pK_a _____Experimental K_a of Acetic Acid _____

(3 pts.)

Percent error of K_a _____

Show both percent error calculations (4 pts.)

Unknown code _____

Volume of base to reach equivalence point _____

pH at equivalence point _____

Volume of base halfway to equivalence point _____

Gsu Chemistry 1212 Lab Manual

Jack Duff, Alan Gabrielli



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CHEM 111 Lab Manual General Chemistry Christofield,2019-08-12 *Laboratory Manual for Principles of General Chemistry* J. A. Beran,Mark Lassiter,2022-09-22 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don t write experiments A student s scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition s experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment s Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

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