

LAB REPORT

1) The two peaks of chlorine are 112 + 114.

molar mass of chlorobenzene = 112, 114

molar mass of benzene = 77

$$112 - 77 = 35$$

$$114 - 77 = 37$$

two isotopes
of Chlorine

The two peaks of bromine are 156 + 158.

$$156 - 77 = 79$$

$$158 - 77 = 81$$

two isotopes
of Bromine

~~The two peaks of bromine are~~

The natural abundances of chlorine - Cl = the lengths of 112 + 114 from the mass spectrum are 8.3 cm and 2.3 cm.

$$^{35}\text{Cl} = \frac{8.3}{8.3 + 2.3} = \frac{8.3}{10.6} \times 100\% = 78\%$$

$$^{37}\text{Cl} = \frac{2.3}{8.3 + 2.3} = \frac{2.3}{10.6} \times 100\% = 22\%$$

The natural abundances of bromine - Br = the lengths of 156 + 158 from the mass spectrum are 6.1 cm and 6.0 cm.

$$^{79}\text{Br} = \frac{6.1}{6.1 + 6.0} = \frac{6.1}{12.1} \times 100\% = 50\%$$

$$^{81}\text{Br} = \frac{6.0}{6.1 + 6.0} = \frac{6.0}{12.1} \times 100\% = 50\%$$

Nomenclature Report

1. The trend I noticed with observing the physical characteristics is when a compound was ionic and metal it formed a crystalline solid, but when the compound had a mixture of molecules or had larger molecules they formed amorphous solids.

Signature

Date

Witness/TA

Date

General Chemistry Laboratory Manual Chm 121

Padhraic Smyth



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