

PART II- REVIEW FINAL EXAM- CHEM 111

Name Key
Last First

1. What is the net-ionic equation for the reaction that occurs when:

a. nitrous acid and sodium hydroxide solutions are mixed.



Answer: $\text{HNO}_2 + \text{OH}^- \rightarrow \text{NO}_2^- + \text{H}_2\text{O}$ net-ionic



b. potassium hydroxide and hydrofluoric acid solutions are mixed.



Answer: $\text{OH}^- + \text{HF} \rightarrow \text{F}^- + \text{H}_2\text{O}$ net-ionic



2. Which of the elements given below will have?

Ba Mg Cl Te I

a. largest atomic size. ☒ a.

b. highest ionization potential

c. lowest electronegativity Ba

d. lowest electron affinity **Ba**



3.a. Name the three classes of compounds that are considered as electrolytes.

L. Schuler, S. Schuler

ii. Strong acidsiii. Strong bases

b. Label each of the following as strong electrolyte, weak electrolyte, or non-electrolyte.

(You must know the list of soluble salts, strong acids and strong bases to answer this question.)

1. $\text{NaC}_2\text{H}_3\text{O}_2$ strong

ii. $\text{HC}_2\text{H}_3\text{O}_2$ 4.74 K

iii. HCN weak

iv. HCHO_2 $\text{Li}^+\text{A}^-\text{K}$ v. NH_4BrO_3 5.1 g/molvi. C_4H_{12} nonelectrolyte

4. Tell if a precipitate may form upon mixing each of the following solutions:

(You must know the list of solubility rules to answer this question)

a. ammonium chloride and lead (II) nitrate

b. barium nitrate and potassium sulfate.

e. nickel (II) sulfate and sodium hydroxide

5. a. Define: isoelectronic

Different species that have the same number and configuration of electrons

b. Which of the following isoelectronic species has the largest size?

Se^{2-}	Br^-	Kr	Sr^{2+}
32 protons 34	35	36	38

on solutions 36	36	36	36
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The higher the nuclear charge (z of protons), the smaller the radius. The shells (energy levels) contract due to the greater attractive force of the nucleus. The radius decreases with increasing atomic number (no. of protons). Se^{2-} will have the largest size.

General Chemistry Ii Final Exam

Yan Bai



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