

## Unit 5: Chemical Kinetics Exam

1.

Trial	$[A_2]$	$[B]$	Initial rate ( $M/s$ )
1	0.10	0.50	$2.5 \times 10^{-3}$
2	0.20	0.50	$5.0 \times 10^{-3}$
3	0.30	0.25	$5.0 \times 10^{-3}$
4	0.30	0.10	$1.0 \times 10^{-3}$

An experiment was conducted to determine the rate law for the reaction  $A_2(g) + B(g) \rightarrow A_2B(g)$ . The table above shows the data collected. Based on the data in the table, which statement is correct?

- (A) Since the rate law can be expressed as **rate** =  $k[A_2]^2$ , tripling the concentration of  $A_2$  will cause a 9-fold increase in the rate of the reaction.
- (B) Since the rate law can be expressed as **rate** =  $k[A_2][B]$ , doubling the concentrations of  $A_2$  and  $B$  will quadruple the rate of the reaction.
- (C) Since the rate law can be expressed as **rate** =  $k[A_2]^2[B]$ , tripling the concentration of  $A_2$  while keeping the concentration of  $B$  constant will triple the rate of the reaction.
- (D) Since the rate law can be expressed as **rate** =  $k[A_2][B]^2$ , doubling the concentration of  $B$  while keeping the concentration of  $A_2$  constant will double the rate of the reaction.

2.  $2\text{NO}(g) + \text{Cl}_2(g) \rightarrow 2\text{NOCl}(g)$ 

Experiment	$[\text{NO}]$	$[\text{Cl}_2]$	Initial Rate of NOCl Formation ( $M/s$ )
1	0.0250	0.0250	$4.55 \times 10^{-3}$
2	0.0250	0.1000	$9.10 \times 10^{-3}$
3	0.0500	0.0250	$1.82 \times 10^{-2}$

The initial rates of the reaction represented by the equation shown above were measured for different initial concentrations of  $\text{NO}(g)$  and  $\text{Cl}_2(g)$ . Based on the data given in the table above, which of the following is the rate law expression for the reaction, and why?

- (A) **Rate** =  $k[\text{NO}]^2$ , because the initial rate quadrupled when  $[\text{NO}]$  was doubled but remained constant when  $[\text{Cl}_2]$  was doubled.
- (B) **Rate** =  $k[\text{NO}][\text{Cl}_2]$ , because the initial rate doubled when either  $[\text{NO}]$  or  $[\text{Cl}_2]$  was doubled.
- (C) **Rate** =  $k[\text{NO}][\text{Cl}_2]^2$ , because the initial rate doubled when  $[\text{NO}]$  was doubled and quadrupled when  $[\text{Cl}_2]$  was doubled.
- (D) **Rate** =  $k[\text{NO}]^2[\text{Cl}_2]$ , because the initial rate quadrupled when  $[\text{NO}]$  was doubled and doubled when  $[\text{Cl}_2]$  was doubled.

# General Chemistry II Chapter Chemical Kinetics Exam

**L Darling-Hammond**



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