

Orbitals and Bonding Test #14 SCH4U Name:

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Assessment of Learning

Multiple Choice: Circle the most correct answer. (15K)

1. After the 5s subshell of an atom is filled with electrons, the next electron added will enter the

- A) 4d subshell B) 5p subshell
C) 5d subshell D) 4f subshell

2. Which of the following is a correct set of quantum numbers for an electron in a 5f orbital?

- A) $n = 5, l = 3, m_l = +1$ B) $n = 5, l = 2, m_l = +3$ C) $n = 4, l = 3, m_l = 0$
D) $n = 4, l = 2, m_l = +1$ E) $n = 5, l = 4, m_l = 3$

3. Which one of the following combinations of quantum numbers is not allowed?

	n	l	m_l
A)	3	0	0
B)	3	1	-1
C)	3	2	2
D)	3	2	-1
E)	3	3	2

4. The orientation in space of an atomic orbital is associated with

- A) the principal quantum number (n). B) the angular momentum quantum number (l).
C) the magnetic quantum number (m_l). D) the spin quantum number (m_s).
E) none of the above

5. Consider the following electron configurations to answer the question that follows:

(i) $1s^2 2s^2 2p^5 3s^1$ (ii) $1s^2 2s^2 2p^4 3s^2$ (iii) $1s^2 2s^2 2p^5 3s^2 3p^1$

(iv) $1s^2 2s^2 2p^5 3s^2 3p^2$ (v) $1s^2 2s^2 2p^5 3s^2 3p^3$

The electron configuration belonging to the atom with the highest 1st ionization energy is

- A) (i) B) (ii) C) (iii) D) (iv) E) (v)

6. Which atomic orbitals overlap to form the C-O bond of dimethyl ether, $(CH_3)_2O$?

- A) C $2s^2$ + O $2s^2$ B) C $2s^2$ + O $2p$
C) C $2s^2$ + O $2s^2$ D) C $2s^2$ + O $2s^2$

General Chemistry Chapter 14 Covalent Bonding Orbitals Exam

David B. Troy, Paul Beringer



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