

1	H	1.008	(1)
2	He	4.003	(18)
3	Li	6.941	(1)
4	Be	9.012	(2)
5	B	10.811	(13)
6	C	12.011	(14)
7	N	14.007	(15)
8	O	15.999	(16)
9	F	18.998	(17)
10	Ne	20.180	(18)
11	Na	22.990	(1)
12	Mg	24.305	(2)
13	Al	26.982	(13)
14	Si	28.086	(14)
15	P	30.974	(15)
16	S	32.06	(16)
17	Cl	35.453	(17)
18	Ar	39.948	(18)
19	K	39.098	(1)
20	Ca	40.078	(2)
21	Sc		(3)
22	Ti	47.88	(4)
23	V	50.942	(5)
24	Cr	51.996	(6)
25	Mn	54.938	(7)
26	Fe	55.845	(8)
27	Co	58.933	(9)
28	Ni	58.693	(10)
29	Cu	63.546	(11)
30	Zn	65.38	(12)
31	Ga	69.723	(13)
32	Ge	72.630	(14)
33	As	74.922	(15)
34	Se	78.96	(16)
35	Br	79.904	(17)
36	Kr	83.80	(18)
37	Rb	85.468	(1)
38	Sr	87.62	(2)
39	Y		(3)
40	Zr	91.224	(4)
41	Nb	92.906	(5)
42	Mo	95.94	(6)
43	Tc		(7)
44	Ru	101.07	(8)
45	Rh	102.91	(9)
46	Pd	106.42	(10)
47	Ag	107.87	(11)
48	Cd	112.41	(12)
49	In	114.82	(13)
50	Sn	118.71	(14)
51	Pb	127.4	(15)
52	Hg	200.59	(16)
53	Tl	204.38	(13)
54	Pb	207.2	(14)
55	Bi	208.98	(15)
56	Po		(16)
57	At		(17)
58	Rn		(18)
59	Fr		(1)
60	Ra		(2)
61	Ac		(3)
62	Th	232.04	(4)
63	Pa	231.04	(5)
64	U	238.03	(6)
65	Np		(7)
66	Pu	244.06	(8)
67	Am		(9)
68	Cm		(10)
69	Bk		(11)
70	Cf		(12)
71	Es		(13)
72	Fm		(14)
73	Mendelevium		(15)
74	Nobelium		(16)
75	Lanthanum	138.91	(13)
76	Ce	140.12	(14)
77	Pr	140.91	(15)
78	Nd	144.24	(16)
79	Pm		(17)
80	Sm	150.36	(18)
81	Eu	151.96	(19)
82	Gd	157.25	(20)
83	Tb	158.93	(21)
84	Dy	162.50	(22)
85	Ho	164.93	(23)
86	Er	167.26	(24)
87	Tm	168.93	(25)
88	Yb	173.05	(26)
89	Lu	174.97	(27)
90	Hafnium	178.49	(4)
91	Tantalum	180.95	(5)
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94	Osmium	190.23	(8)
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104	Radon		(18)
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132	Flerovium		(28)
133	Moscovium		(29)
134	Livermorium		(30)
135	Tennessine		(31)
136	Oganesson		(32)

Nonmetals

The I–II group designation has been recommended by the International Union of Pure and Applied Chemistry (IUPAC) but is not yet in wide use. In this text we use the standard IUPAC notation for group numbers (1A–8A and 1B–10B). No numbers have been assigned for elements 11–18. In 2011 IUPAC revised the atomic numbers of these elements. The changes are minor and they are not adopted in the present edition of this text.

89PA-TV84-QVBQ-3GNT-V3TB

General Chemistry

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**Stephen J. Fonash, Marcel Van de
Voorde**



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Pengantar Kimia dan Metode Ilmiah Lulu Nurdini, Faradisa Anindita, Santi Febri Arianti, Syarifuddin Syarifuddin, Fakhrotun Nisa, Nadya Syarifatul Fajriyah, Matius Stefanus Batu, Maria Assumpta Nogo Ole, Mayshah Purnamasari, IGP Gede Rumayasa Yudana, 2025-06-18 Buku Pengantar Kimia dan Metode Ilmiah adalah panduan komprehensif yang dirancang untuk memperkenalkan dasar dasar ilmu kimia kepada pembaca pemula terutama mahasiswa tahun pertama Buku ini disusun dengan pendekatan yang sistematis dan aplikatif menjembatani pemahaman antara teori kimia dasar dan praktik ilmiah yang sesungguhnya Konten buku ini mencakup pengenalan terhadap hakikat ilmu kimia struktur dan tata nama senyawa reaksi kimia hukum dasar kimia hingga penerapan metode ilmiah dalam konteks kimia Pembaca akan diajak memahami konsep melalui contoh contoh yang dekat dengan kehidupan sehari hari sekaligus dilatih untuk berpikir kritis dan analitis sebagaimana dituntut dalam dunia sains Buku ini tidak hanya menjelaskan konsep kimia tetapi juga mengapa dan bagaimana yang menjadikannya sumber belajar yang tidak hanya informatif tetapi juga reflektif dan aplikatif Dengan demikian buku ini sangat tepat digunakan sebagai pengantar dalam perkuliahan pelatihan maupun studi mandiri dalam bidang kimia dan sains pada umumnya

CMS Cara Menguasai Soal Kimia SMA dan MA Latihan Soal dan Pembahasan Plus HOTS Enik Suyahni, Buku ini terdiri atas 12 bab yang mencakup seluruh materi pada Kimia SMA MA dari kelas X sampai kelas XII Pada setiap bab terdapat soal latihan dan pembahasannya soal pemantapan dan variasi soal

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