- According to the Bronsted—Lowry definition, which chemical species can function both as an acid and as a base?
- 17 a CI so,2

17

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- c. NH.
- d. HCO₃
- e. H₃O*
- In the reaction CN⁻ + H₂O ←→ HCN + OH⁻ which is an acid—base conjugate pair?
 - a. H2O and HCN
 - b. H₂O and OH
 - c. CN⁻ and H₂O
 - d. HCN and OH
- Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX
 - a. It is less basic than a 1 M solution of NaY.
 - It is more basic than a 1 M solution of NaY.
 - It yields a neutral solution.
 - It is more concentrated than a 1 M solution of NaY.
- 4. The pH of a 0.03 M HCl solution is
- a. 1.5
 - b. 2.5
 - c. 3.5 d. 12.5
 - 5. The pH of a 1.0 x 10⁻³ M Ba(OH)₃ solution at
 - 25°C is a. 2.7
 - a. 2.7 A. 3.0
 - c. 11.0
 - d. 11.3
 - 6. Which is the strongest acid?
- a. HCiO
 - & HClO₂
 - c. HClO₃
 - d. HClO₂
 - 7. Which salt reacts with water (hydrolyzes) to produce a basic solution?
- F & NaC2H3O2
 - A. NaNO
 - c. NH₄Cl
 - at BaSO.

8. The weakest of the bases listed is

Aeid	Conjugate Base	Ka (Ionization Constant of Acid)	
HCI	CI	100% ionized	
HSO ₄	SO,2	1.2 x 10 ⁻²	
H ₂ S	HS	5.7x10 ⁻⁸	
HIS	S	1.2×10 ⁻¹⁵	7

at CE

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- & CN
- c. HS
- d. 52-
- e. SO₂2-

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Table 1-1	100			-			7	-	



- Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate KC₂H₃O₂, is less alkaline than a 0.10 M solution of potassium cyanide, KCN?
- A. Hydrocyanic acid is a weaker acid than acetic acid.
 - Hydrocyanic acid is less soluble in water than acetic acid.
 - Cyanides are less soluble than acetates
 - A Acetic soid is a weaker soid than hydrocyanic soid
 - In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of
 - a. a 0.100 M sodium solution.
 - a 0.0500 M sodium hydroxide solution.
 - c. a 0.0500 M benzoic acid solution.
 - d. a 0.0500 M sodium benzoate solution.

General Chemistry 1 Acs Practice Test

JA Banks

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