

Which of these are **true** and which are **false** about acid-base reactions?

- In general, an acid-base reaction will be favored when a weaker acid and a weaker base combine to form a stronger acid and a stronger base
Other way around: will be favored when a stronger acid and a stronger base combine to give a weaker acid and a weaker base.
T F
- Generally, strong acids have weak conjugate bases
True - and weak acids have strong conjugate bases
T F
- Strong acids tend to have high pK_a values
Strong acids have low (strongly negative) pK_a values - e.g. HCl has a pK_a of -8
T F
- The greater the acidity of a molecule, the weaker a base it is
Not quite correct. The stronger the acid, the weaker is its conjugate base.
T F
- The inductive effect of electronegative atoms near a carboxylic acid group increases the acidity of the compound.
T F
- Any factor which stabilizes negative charge generally tends to decrease the acidity of an acid
False - increases acidity, not decreases.
T F
- The pK_a of NH_4 is 10 and the pK_a of H_2O is 14, which means that H_2O is a stronger acid than H_2O
False - NH_4 is a stronger acid by a factor of 4 pK_a units (about 10^4)
T F
- The acidity of a molecule can be increased by resonance stabilization of its conjugate base.
T F
- sp -hybridized C-H bonds are more acidic than sp^3 -hybridized C-H bonds since the resulting negative charge is closer to the nucleus
T F
- Acid-base reactions where a stronger base is formed from a weaker base never occur.
They do occur and they can be important. E.g. NH_3 can deprotonate H_2O even though it is a weaker base by a factor of 10^4 compared to HO^- . Even a difference in 9-10 pK_a units can be significant for moving some reactions forward.
T F

General Chemistry Acid Base Practice Problems

Marco Casella



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