

TOPICS:

- ❖ Intermolecular Forces
- ❖ Colligative Properties

Intermolecular Forces

→ The attraction between molecules

Intermolecular forces are always due to coulombic attraction, which is the attraction between molecules due to electric charge



Ion -Dipole

- Ion and polar molecule interact (Polar; unequal pull between electrons, often has lone pairs on central atoms.)

Hydrogen Bonding

- Looking for bonds between hydrogen and fluorine, oxygen, and nitrogen

Dipole-Dipole

- Polar molecules

London Dispersion/Van der Waals

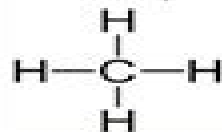
- Mainly nonpolar molecules

Arrange in increasing order: CH_4 , H_2S , CCl_4 , NH_3

Step 1 - draw lewis structures

Step 2 - determine intermolecular forces

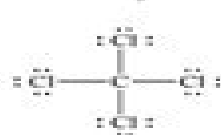
london dispersion



dipole-dipole



london dispersion



hydrogen bonding



Increasing Order: CH_4 , CCl_4 , H_2S , NH_3

*The larger the molar mass, the stronger the intermolecular force. Use this as a tiebreaker

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