

09/08/2020

Chapter 1: General Chem Review - Part 1

- 1.1 Reactions: a) molecules collide
b) bonds broken/formed

Organic compounds \rightarrow compounds w carbon atoms

- 1.2 Structural Theory: substances are defined by a specific arrangement
compounds differ based on the way the atoms are bonded

Constitutional isomer: same molecular formula but different structures

Atoms commonly bonded to C: N, O, H & halides

tetravalent
(4 bonds)

trivalent
(3 bonds)

divalent
(2 bonds)

monovalent
(one bond)

- 1.3 Covalent bond: pair of e^- shared by two atoms

Energy released when forming bond \rightarrow equilibrium internuclear distance

internuclear repulsion if contracted further \rightarrow (bond length)

$\uparrow \uparrow$ space $= \downarrow$ energy of electrons

energy @ equilibrium
is bond strength

Potential Energy & stability: \downarrow PE = most stable

What keeps bond @ optimal length?

+ nuclei & - electron \rightarrow attractive forces

+ nuclei & + nuclei \rightarrow repulsive forces \leftarrow - electron & - electron

Lewis Structure Examples:



- 1.4 Formal charge: valence electrons it has vs. valence electrons need to be neutral
eg. CH_3O



C has 4 valence electrons \rightarrow no formal charge
Needs 4 more to be neutral

O has 7 electrons \rightarrow 1- formal charge
Needs 1 less to be neutral

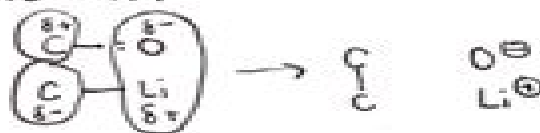
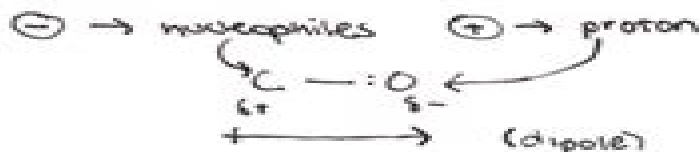
- 1.5 Electronegativity: increase L to R & B to T
F is most EN

\uparrow carbon usually forms polar covalent bonds

Covalent bond: EN difference < 0.5

Polar covalent bond: EN difference $0.5 - 1.7$

Ionic bond: EN difference > 1.7



Gen Chem 09 10 Sem 1 Review Answers

Michal Rosen-Zvi



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Survival Guide to Organic Chemistry Patrick E. McMahon, Bohdan B. Khomtchouk, Claes Wahlestedt, 2016-12-19 The Survival Guide to Organic Chemistry Bridging the Gap from General Chemistry enables organic chemistry students to bridge the gap between general chemistry and organic chemistry It makes sense of the myriad of in depth concepts of organic chemistry without overwhelming them in the necessary detail often given in a complete organic chemistry text Here the topics covered span the entire standard organic chemistry curriculum The authors describe subjects which require further explanation offer alternate viewpoints for understanding and provide hands on practical problems and solutions to help master the material This text ultimately allows students to apply key ideas from their general chemistry curriculum to key concepts in organic chemistry Key Features Reviews key general chemistry concepts and techniques adapted for application to important organic principles Provides practical guidance to help students make the notoriously well known and arduous transition from general chemistry to organic chemistry Explains organic concepts and reaction mechanisms generally expanding the focus on how to understand each step from a more intuitive viewpoint Covers concepts that need further explanation as well as those that summarize and emphasize key ideas or skills necessary in this field An added bonus is help with organizing principles to make sense of a wide range of similar reactions and mechanisms Implements a user friendly process to achieve the end result of problem solving Covers organic chemistry I and II concepts at the level and depth of a standard ACS organic chemistry curriculum features practice problems and solutions to help master the material including an extensive and comprehensive bank of practice exams with solutions *Science Tests and Reviews* Buros Center, 1975 Science Tests and Reviews consisting of science sections of the first seven MMYs and Tests in Print II includes 217 original test reviews written by 81 specialists 18 excerpted test reviews 270 references on the construction use and validity of specific tests a bibliography on in print science tests references for specific tests cumulative name indexes for specific tests with references a publishers directory title index name index and a scanning index The 97 tests covered fall into the following categories 23 general 14 biology 35 chemistry 3 geology 6 miscellaneous and 16 physics **Proceedings of NECC 1979, National Educational Computing Conference** Diana Harris, 1979 **Cumulated Index Medicus**, 1965

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baccalaureate degree This essential resource offers pertinent statistical data on such topics as tuition room and board admission requirements financial aid enrollments student life library holdings accelerated and study abroad programs departments and teaching staff buildings and grounds and degrees conferred Volume two of the set provides four indexes including an institutional Index a subject accreditation index a levels of degrees offered index and a tabular index of summary data by state These helpful indexes allow readers to find information easily and to make comparisons among institutions effectively Also contained within the text are charts and tables that provide easy access to comparative data on relevant topics

HOW TO LEARN AND TEACH IN MEDICAL SCHOOL Mark E. Quirk, 1994-01-01 This book can be used by students faculty and administrators to enhance the process of medical education Medical students can use it to develop and refine their learning skills assess their individual learning needs identify resources to meet these needs and to prepare themselves for critical learning events and transitions Faculty can use this book to understand the learning process and develop teaching skills thereby providing a framework for critically assessing how they teach Administrators can use it to expose inadequacies in the medical education process and to strengthen the teaching and learning environment Certain chapters of this book could serve as required readings for a medical school course on learning The development and implementation of such a course and the commensurate application of the espoused principles of learning and teaching to the educational process would signal true medical educational reform It would signify that critical attention is being paid to those aspects of learning medicine thought to be crucial reading the literature self directedness critical thinking problem solving lifelong learning and creativity Such a course would be as important as any basic science course or clinical rotation to the training of future physicians who would use these skills throughout their professional lives The book could serve as a foundation for faculty development with medical educators In this regard Part One provides insight into how medical students learn and has significant implications for teaching Parts Two and Three provide practical guidelines for teaching and curriculum development from a learner centered perspective

El-Hi Textbooks in Print, 1977

A Study of High School Seniors of Superior Ability Emmett Lee Schott, 1926

Psychology of Academic Cheating Eric M. Anderman, Tamera B. Murdock, 2011-04-28 Who cheats and why How do they cheat What are the consequences What are the ways of stopping it before it starts These questions and more are answered in this research based investigation into the nature and circumstances of Academic Cheating Cheating has always been a problem in academic settings and with advances in technology camera cell phones the internet and more pressure than ever for students to test well and get into top rated schools cheating has become epidemic At the same time it has been argued the moral fiber of society as a whole has dampened to find cheating less villainous than it was once regarded Who cheats Why do they cheat and Under what circumstances Psychology of Academic Cheating looks at personality variables of those likely to cheat but also the circumstances that make one more likely than not to try cheating Research on the motivational aspects of cheating and what

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