



Unit 6 Test Study Guide

January 13th, 2022

Brensted-Lowry Acid/Base Theory

- Acids: Proton (H^+) donor
- Bases: Proton (H^+) acceptor
- Conjugate Acids-Base Pair: Two substances that differ only by the presence of an H^+
 - Conjugate Acid: Formed when a proton is added to a base (positive charge)
 - Add a Hydrogen to the formula and add +1 charge
 - Conjugate Base: Formed when a proton is removed from an acid (negative charge)
 - Take away a Hydrogen from the formula and subtract -1 charge
- Amphoteric: The ability for a compound to act as an acid OR as a base (EX: H_2O)
- Acids are classified based on how many H^+ are present
 - Monoprotic: 1 H^+
 - Diprotic: 2 H^+
 - Triprotic: 3 H^+
 - Diprotic and triprotic acids are also amphoteric because they have more than one proton that they can lose

Properties of Strong and Weak Acids/Bases

| Strong Acids/Bases | Weak Acids/Bases |
|---|--|
| <ul style="list-style-type: none">• Strong electrolytes• Ionize completely in water• Equilibrium is not present<ul style="list-style-type: none">◦ Very unlikely that the ions reform into products• Very weak attraction between ions<ul style="list-style-type: none">◦ Split apart easily | <ul style="list-style-type: none">• Weak electrolytes• Does not ionize completely in water• Equilibrium exists• Strong attraction between ions<ul style="list-style-type: none">◦ Do not split apart easily |
| <p style="text-align: center;">Strong acid</p>  | <p style="text-align: center;">Weak acid</p>  |

Equilibrium And Acid Base Study Guide

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Equilibrium And Acid Base Study Guide:

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