

# Enthalpy Worksheet

## Chemistry II

using the average bond energy table below, calculate the enthalpy change in each of the following reactions.

Step 1. Balance the reaction.

Step 2. Classify the reaction.

Step 3. Determine bonds broken and calculate the positive enthalpy change of the reactants. For more complex molecules, draw their structure to help determine the bonds.

(NOTE - if two compounds contain the same types of bonds, they may be combined)

Step 4. Determine the bonds formed and calculate the negative enthalpy change of the products.

Step 5. Find the total enthalpy change by subtracting that of the products from the reactants.

H-Br 432 kJ/mol	I-I 148 kJ/mol	C-Cl 339 kJ/mol	C=O 1072 kJ/mol
H-Cl 295 kJ/mol	O-Br 467 kJ/mol	N-Br 291 kJ/mol	C-C 347 kJ/mol
C-H 413 kJ/mol	Cl-Cl 239 kJ/mol	O=O 495 kJ/mol	C=O 745 kJ/mol
C=O 558 kJ/mol	C=C 614 kJ/mol	N-C 293 kJ/mol	C=O (for CO(g)) 799 kJ/mol
H-Br 348 kJ/mol	Br-Br 193 kJ/mol		

EXAMPLE:



	Bonds Broken			Bonds made	
	H <sub>2</sub>	Br <sub>2</sub>	→	HBr	
Bonds	1 H-H	1 Br-Br	→	2 H-Br	
Bond Energies	432	193		2(368)	
Total	625		-	736	= -111 kJ/mol



	Bonds Broken			Bonds made	
	H <sub>2</sub>	I <sub>2</sub>	→	HI	
Bonds	1 H-H	1 I-I	→	2 H-I	
Bond Energies	432	148		2(295)	
Total	580		-	590	= -10 kJ/mol



	Bonds Broken				Bonds made	
	C <sub>2</sub> H <sub>4</sub>			→	CO <sub>2</sub>	H <sub>2</sub> O
Bonds	1 C=C	4 C-H	1 O=O	→	2 C=O	2 O-H
Bond Energies	1316	4(413)	4(495)		2(799)	2(467)
Total	4858			-	2322	= 2536 kJ/mol



# Enthalpy Change Chemistry Answer Sheet

**Xiaolong Qi**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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