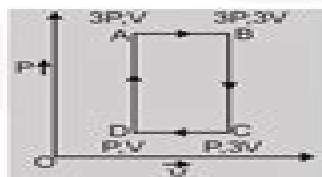


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- Pressure-temperature relationship for an ideal gas undergoing adiabatic change is ($\gamma = C_p/C_v$)
 (1) $pT^\gamma = \text{constant}$ [AIPMT 92, HARYANA PMT 01, RPET 2001]
 (2) $pT^{1-\gamma} = \text{constant}$
 (3) $pT^{1+\gamma} = \text{constant}$
 (4) $p^{1-\gamma}T^\gamma = \text{constant}$
- An inventor claims that when temperatures of source and sink are 127°C and 27°C respectively then efficiency of engine is 26%, then :- [AIPMT 2001]
 (1) It is impossible (2) It is possible but possibility is less
 (3) It's possibility is high (4) Data is insufficient
- If the system takes 100 cal. heat, and releases 80 cal to sink, if source temperature is 127°C find the sink temp. :- [AIIMS 2001]
 (1) 47°C (2) 127°C (3) 67°C (4) 107°C
- The efficiency of carnot engine is 50% and temperature of sink is 500K. If temperature of source is kept constant and its efficiency raised to 60%, then the required temperature of the sink will be:- [MANIPAL 2001, AIPMT 2002]
 (1) 100 K (2) 600 K (3) 400 K (4) 500 K
- An ideal gas heat engine operates in carnot cycle between 227°C and 127°C . It absorbs 6×10^4 cal of heat at higher temperature. Amount of heat converted to work is. [AIPMT 2005]
 (1) 2.4×10^4 cal (2) 6×10^4 cal (3) 1.2×10^4 cal (4) 4.8×10^4 cal
- An ideal monoatomic gas is taken round the cycle ABCDA as shown in following P - V diagram. The work done during the cycle is : [RPET 92, 2000, IIT 83, AIPMT 90, 92, 2007]



- (1) PV (2) $2 PV$ (3) $4 PV$ (4) Zero
- The $\left(\frac{W}{Q}\right)$ of a carnot-engine is $\frac{1}{6}$, now the temp. of sink is reduced by 62°C , then this ratio becomes twice, therefore the initial temp. of the sink and source are respectively :- [AIPMT 2007]
 (1) 33°C , 67°C (2) 37°C , 99°C
 (3) 67°C , 33°C (4) 97 K, 37 K

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Table of Contents Me121thermodynamics Previous Year Question Paper

1. Understanding the eBook Me121thermodynamics Previous Year Question Paper
 - The Rise of Digital Reading Me121thermodynamics Previous Year Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Me121thermodynamics Previous Year Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Me121thermodynamics Previous Year Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Me121thermodynamics Previous Year Question Paper
 - Personalized Recommendations
 - Me121thermodynamics Previous Year Question Paper User Reviews and Ratings
 - Me121thermodynamics Previous Year Question Paper and Bestseller Lists
5. Accessing Me121thermodynamics Previous Year Question Paper Free and Paid eBooks
 - Me121thermodynamics Previous Year Question Paper Public Domain eBooks
 - Me121thermodynamics Previous Year Question Paper eBook Subscription Services
 - Me121thermodynamics Previous Year Question Paper Budget-Friendly Options
6. Navigating Me121thermodynamics Previous Year Question Paper eBook Formats

- ePub, PDF, MOBI, and More
- Me121thermodynamics Previous Year Question Paper Compatibility with Devices
- Me121thermodynamics Previous Year Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Me121thermodynamics Previous Year Question Paper
 - Highlighting and Note-Taking Me121thermodynamics Previous Year Question Paper
 - Interactive Elements Me121thermodynamics Previous Year Question Paper
- 8. Staying Engaged with Me121thermodynamics Previous Year Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Me121thermodynamics Previous Year Question Paper
- 9. Balancing eBooks and Physical Books Me121thermodynamics Previous Year Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Me121thermodynamics Previous Year Question Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Me121thermodynamics Previous Year Question Paper
 - Setting Reading Goals Me121thermodynamics Previous Year Question Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Me121thermodynamics Previous Year Question Paper
 - Fact-Checking eBook Content of Me121thermodynamics Previous Year Question Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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