

Name: _____ Date: _____ Period: _____

Midterm Exam Study Guide

d. Isotopes

- i. Isotopes differ in their _____ and number of _____

e. Average atomic mass Formula:

f. Ions

- i. _____ have lost electrons and have a _____ charge
- ii. _____ have gained electrons and have a _____ charge

g. Group Numbers be able to locate them

- i. Alkali Metals
- ii. Alkaline earth Metals
- iii. Transition Metals
- iv. Inner Transition Metals
- v. Halogens
- vi. Noble Gases

h. Periodic Trends Give your notes for reference (the charts for each of these)

- i. Atomic Radius/Size
- ii. Ionic Radius/Size
- iii. Ionization Energy
- iv. Electronegativity

i. Valence Electrons be able to identify # of valence electrons and charge for each group

- i. Lewis-Dot structures

F. Nuclear Decay and Half Life

- a. ${}^4_2\text{He}$ or ${}^4_2\text{He}$ is called an _____ particle which has the weakest radiation.
- b. ${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$ is called an _____ particle which has moderate radiation and is an electron.
- c. ${}^4_{-2}\text{He}$ is called an _____ particle which has the strongest radiation and highest energy.
- d. Half-life
 - i. Definition: _____
 - ii. Know how to read a graph and calculate half-life

d. Electron Configurations

- a. Be able to draw them/identify them from another student
- b. Hund's Rule
- c. Pauli Exclusion Principle
- d. Aufbau Principle

e. Ionic Compounds

- a. Ionic compounds are made of _____ + _____
- b. Their overall charge is _____
- c. Know how to use Lewis-Dot Structures to draw bonds between cation and anion

Mcps Chemistry Semester Review Guide Answers

Zhe Xu



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