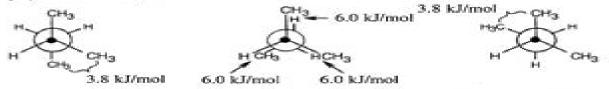
3.42

The energy difference between the two conformations is (11.0 + 6.0 + 4.0) kJ/mol – 3.8 kJ/mol = 17.2 kJ/mol.

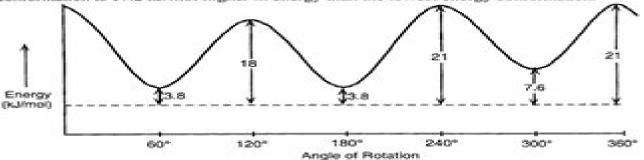
(c) Consider the least stable conformation to be at zero degrees. Keeping the front of the projection unchanged, rotate the back by 60° to obtain each conformation.



at 120°; energy = 3.8 kJ/mol at 120°; energy = 18.0 kJ/mol at 180°; energy = 3.8 kJ/mol



Use the lowest energy conformation as the energy minimum. The highest energy conformation is 17.2 kJ/mol higher in energy than the lowest energy conformation.



Mcmurry Organic Chemistry Solutions Manual Scribd

Francis Carey, Robert Giuliano

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