

## Acid and Base Worksheet - Answers

- 1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair:
- a)  $\text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^-$   
 **$\text{HNO}_3$  and  $\text{NO}_3^-$  make one pair  $\text{OH}^-$  and  $\text{H}_2\text{O}$  make the other**
- b)  $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{NH}_3^+ + \text{OH}^-$   
 **$\text{CH}_3\text{NH}_2$  and  $\text{CH}_3\text{NH}_3^+$  make one pair  $\text{OH}^-$  and  $\text{H}_2\text{O}$  make the other**
- c)  $\text{OH}^- + \text{HPO}_4^{2-} \rightarrow \text{H}_2\text{O} + \text{PO}_4^{3-}$   
 **$\text{HPO}_4^{2-}$  and  $\text{PO}_4^{3-}$  make one pair  $\text{OH}^-$  and  $\text{H}_2\text{O}$  make the other**
- 2) The compound NaOH is a base by all three of the theories we discussed in class. However, each of the three theories describes what a base is in different terms. Use your knowledge of these three theories to describe NaOH as an Arrhenius base, a Brønsted-Lowry base, and a Lewis base.
- **NaOH is an Arrhenius base because it creates  $\text{OH}^-$  ions when placed in water.**
  - **NaOH is a Brønsted-Lowry base because it accepts  $\text{H}^+$  ions from acids.**
  - **NaOH is a Lewis base because the lone pairs on the hydroxide ion can be donated to other compounds.**
- 3) When hydrogen chloride reacts with ammonia, ammonium chloride is formed. Write the equation for this process, and indicate which of the reagents is the Lewis acid and which is the Lewis base.
- $\text{HCl} + \text{NH}_3 \rightarrow \text{NH}_4\text{Cl}$   
 **$\text{NH}_3$  is a Lewis base because it uses its lone pair electrons to pull a hydrogen atom from hydrochloric acid.**  
 **$\text{HCl}$  is a Lewis acid because it accepts electrons from  $\text{NH}_3$  when the H is transferred.**
- 4) Write an equation for each of the following reactions and indicate what each reaction produces: (For each of the following oxides, indicate if it is an acid or basic anhydride).
- a. Carbon dioxide and water  
 $\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3$   
**Acid anhydride**
- b. Sodium oxide and water  
 $\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{NaOH(aq)}$   
**Basic anhydride**
- c. Dichlorine trioxide and water  
 $\text{Cl}_2\text{O}_3 + \text{H}_2\text{O} \rightarrow 2\text{HClO}_2$   
**Acidic anhydride**
- d. Beryllium oxide and water  
 $\text{BeO} + \text{H}_2\text{O} \rightarrow \text{Be(OH)}_2$   
**Basic anhydride**

# Introduction To Acids And Bases Answer Key

**M Woodhall**



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## **Table of Contents Introduction To Acids And Bases Answer Key**

1. Understanding the eBook Introduction To Acids And Bases Answer Key
  - The Rise of Digital Reading Introduction To Acids And Bases Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Introduction To Acids And Bases Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Introduction To Acids And Bases Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Introduction To Acids And Bases Answer Key
  - Personalized Recommendations
  - Introduction To Acids And Bases Answer Key User Reviews and Ratings

- Introduction To Acids And Bases Answer Key and Bestseller Lists
- 5. Accessing Introduction To Acids And Bases Answer Key Free and Paid eBooks
  - Introduction To Acids And Bases Answer Key Public Domain eBooks
  - Introduction To Acids And Bases Answer Key eBook Subscription Services
  - Introduction To Acids And Bases Answer Key Budget-Friendly Options
- 6. Navigating Introduction To Acids And Bases Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Introduction To Acids And Bases Answer Key Compatibility with Devices
  - Introduction To Acids And Bases Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Introduction To Acids And Bases Answer Key
  - Highlighting and Note-Taking Introduction To Acids And Bases Answer Key
  - Interactive Elements Introduction To Acids And Bases Answer Key
- 8. Staying Engaged with Introduction To Acids And Bases Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Introduction To Acids And Bases Answer Key
- 9. Balancing eBooks and Physical Books Introduction To Acids And Bases Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Introduction To Acids And Bases Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Introduction To Acids And Bases Answer Key
  - Setting Reading Goals Introduction To Acids And Bases Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Introduction To Acids And Bases Answer Key
  - Fact-Checking eBook Content of Introduction To Acids And Bases Answer Key
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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