

**Example:** Let us take an example of the manufacturing of ammonia. When one mole of nitrogen gas reacts with three moles of hydrogen it produces two moles of ammonia gas. This is known as 'Forward reaction'.



In contrast, two moles of ammonia gas may also be converted into one mole of nitrogen and three moles of hydrogen, this is known as a 'Reverse reaction'.



When both of these reactions are written together as a reversible reaction they are represented as:



**Q.6:** Explain Dynamic Equilibrium with an example.

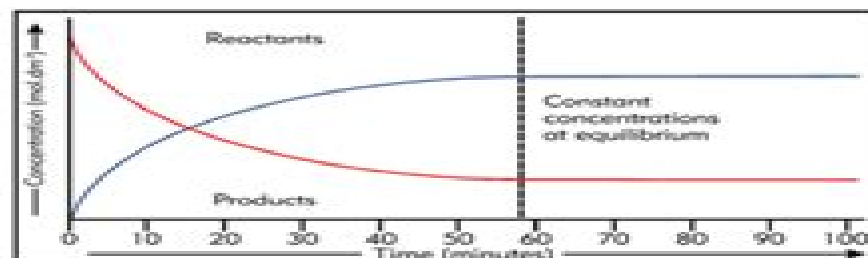
**Ans:** **Dynamic Equilibrium:**

The reaction rate depends on the concentration of the reactants. In the beginning, the quantity of reactant is higher, and the rate of product formation is higher. As the reactant amount decreases, the rate of reactant transformation also decreases, and the rate of product formation decreases. After a certain time, the concentrations of reactants and products become constant, and this state is called dynamic equilibrium.

$$\text{Rate of the forward reaction} = \text{Rate of the reverse reaction}$$

In a reversible reaction, dynamic equilibrium is established before the completion of the reaction. The rate of both forward and reverse reactions becomes equal upon reaching the equilibrium point. The following graph which is of concentrations vs. time shows that the concentrations of both reactants and products become constant at equilibrium.

**Example:** An example of a reaction at equilibrium is a reaction of hydrogen and iodine in a closed container to produce hydrogen iodide. At the start of the reaction, there is a high concentration of hydrogen and iodine and, after that, the concentration decreases as hydrogen iodide is formed. The concentration of hydrogen iodide increases as the forward reaction proceeds. As hydrogen iodide is formed the reverse reaction is then able to occur.



# Multiple Choice Questions And Answers In Chemical Equilibrium

**VM Jensen**



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