

PHYSIQUE-CHIMIE

Calculatrices autorisées. Les parties I et II sont indépendantes.

Partie I - L'humidité dans l'air

La vapeur d'eau désigne la phase gazeuse de l'eau, composé chimique présent sur Terre sous forme solide, liquide ou gazeuse.

Dans tout le problème nous désignons par :

- « air sec », un mélange gazeux de dioxygène et de diazote ;
- « air humide », un mélange gazeux de dioxygène, de diazote et de vapeur d'eau ;
- « air humide saturant », un mélange gazeux de dioxygène, de diazote et de vapeur d'eau en équilibre avec de l'eau liquide.

L'air, sec ou humide, en dehors des conditions de saturation, se comporte comme un mélange idéal de gaz parfaits. Chacun des constituants y est décrit par sa pression partielle p ou sa fraction massique x (ou encore sa fraction molaire w). Les conditions normales de pression correspondent à $p_{atm} = p^{\circ} = 1 \text{ bar}$. On rappelle qu'une température de 0°C équivaut à $273,16 \text{ K}$.

Si l'on fait subir des transformations thermodynamiques à une masse d'air, en évitant de provoquer un quelconque phénomène de condensation, les fractions molaires des gaz restent constantes. Dans ce problème, on admet que seule l'eau est susceptible de se condenser. En conditions saturantes, on admet que la pression partielle en eau est uniquement fonction de la température de l'air. On la note alors $p_{H_2O}^{sat}(T)$.

On appelle température de rosée, la température à laquelle il faut refroidir, à pression constante, une masse d'air humide pour atteindre la saturation, c'est-à-dire l'apparition de la première goutte d'eau liquide.

On appelle pression de vapeur saturante à la température T , notée $p_{H_2O}^{sat}(T)$, la valeur maximale que peut atteindre la pression partielle de vapeur d'eau au sein d'un air humide.

On appelle humidité absolue, la masse de vapeur d'eau présente dans un kilogramme d'air humide.

Meacutemofiches Concours Kineacute Physique Chimie

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