



10. Which of the following are endothermic processes?

- A. ice melting
- B. water condensing
- C. freezing
- D. sublimation

11. Heat absorbed per mole at STP, positive temperature change

12. Which of the following is a statement of the first law of thermodynamics?

- A. the negative change in enthalpy corresponds to an exothermic process
- B. energy lost by the system must be gained by the surroundings, neither of which are created
- C. temperature is a measure of average kinetic energy
- D. enthalpy is a state function

13. The specific enthalpy of combustion of propane, ΔH is negative. This means that the product is _____

- A. a hydrocarbon with the same enthalpy as the reactant
- B. a hydrocarbon and the temperature of the surroundings will decrease
- C. a hydrocarbon and the surroundings will increase
- D. a hydrocarbon and the temperature of the surroundings will increase
- E. none of these is a true statement

14. The reaction



is _____ and therefore heat is _____ from the system.

- A. endothermic, released
- B. exothermic, absorbed
- C. endothermic, released
- D. exothermic, absorbed
- E. thermodynamic value released and absorbed

15. Heat evolved when 1 mole of C_2H_2 is burned with O_2 at STP is -1300 kJ . The reaction is



- A. C_2H_2
- B. O_2
- C. CO_2
- D. H_2O
- E. H_2



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