

Activity B:**Methane reaction**Get the Gizmo ready:

- Select $C + H_2$ becomes CH_4 .



Goal: Predict the results of a chemical reaction.

1. Count: CH_4 is the chemical formula for methane. In order to produce a single molecule of methane, how many carbon (C) atoms are needed? Hydrogen atoms?

2. Predict: Set the number of C atoms to 14 and the number of H_2 molecules to 10.

A. How many total hydrogen atoms are there?

B. How many methane molecules could form from these reactants?

C. After the atoms react, which reactant will be left over?

D. Which reactant will be the limiting reactant?

E. Click **Play** and wait until **Reaction complete** is shown. What happened?

Each hydrogen molecule formed a connection with carbon (4 hydrogen molecules per carbon). There are extra carbon atoms that did not link because there are additional carbon atoms.

3. Test: Click **Reset**. Set the number of C atoms to seven.

A. How many hydrogen molecules (H_2) will be needed for there to be no excess reactants? Use the slider to set the number of H_2 molecules to this value.

B. How many methane molecules will be formed?

C. Click **Play**. What happened?

Since all of the hydrogen molecules linked to the carbon, there were no extra reactants.

4. Solve: For each **chemical equation** below, write the number of product molecules that will form from the reaction. Then, circle the limiting reactant. (Note: The **coefficients** in front of the reactants indicate the number of reactant molecules or atoms present.)



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Lingsheng Yao



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Table of Contents Full Version Student Exploration Limiting Reactants Answer Key Gizmo

1. Understanding the eBook Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - The Rise of Digital Reading Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - Advantages of eBooks Over Traditional Books
2. Identifying Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - User-Friendly Interface
4. Exploring eBook Recommendations from Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - Personalized Recommendations

- Full Version Student Exploration Limiting Reactants Answer Key Gizmo User Reviews and Ratings
- Full Version Student Exploration Limiting Reactants Answer Key Gizmo and Bestseller Lists
- 5. Accessing Full Version Student Exploration Limiting Reactants Answer Key Gizmo Free and Paid eBooks
 - Full Version Student Exploration Limiting Reactants Answer Key Gizmo Public Domain eBooks
 - Full Version Student Exploration Limiting Reactants Answer Key Gizmo eBook Subscription Services
 - Full Version Student Exploration Limiting Reactants Answer Key Gizmo Budget-Friendly Options
- 6. Navigating Full Version Student Exploration Limiting Reactants Answer Key Gizmo eBook Formats
 - ePub, PDF, MOBI, and More
 - Full Version Student Exploration Limiting Reactants Answer Key Gizmo Compatibility with Devices
 - Full Version Student Exploration Limiting Reactants Answer Key Gizmo Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - Highlighting and Note-Taking Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - Interactive Elements Full Version Student Exploration Limiting Reactants Answer Key Gizmo
- 8. Staying Engaged with Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Full Version Student Exploration Limiting Reactants Answer Key Gizmo
- 9. Balancing eBooks and Physical Books Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Full Version Student Exploration Limiting Reactants Answer Key Gizmo
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - Setting Reading Goals Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Full Version Student Exploration Limiting Reactants Answer Key Gizmo
 - Fact-Checking eBook Content of Full Version Student Exploration Limiting Reactants Answer Key Gizmo

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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