

## Gas Stoichiometry: Multiple Choice

1. What volume of pure oxygen gas,  $O_2$ , measured at 546 K and 101.325 kPa is formed by the complete decomposition of 0.200 mol  $Ag_2O$  according to the following equation:



- A) 1.12 L      B) 4.48 L      C) 2.24 L      D) 8.96 L      E) 17.82 L

ANS: B

1. How many moles of  $KClO_3$  are needed to form 2.8 L of  $O_2$ , measured at STP, according to the following reaction:  $2KClO_3 \rightarrow 2KCl + 3O_2$

- A) 0.083 mol    B) 0.25 mol    C) 0.17 mol    D) 0.50 mol    E) 0.33 mol

ANS: A

1. What volume of chlorine gas, measured at 0°C and 101.325 kPa, is required to react completely with 5.61 g of KOH in the following reaction:  $Cl_{2(g)} + 2KOH \rightarrow KCl + KClO + H_{2O}$

- A) 0.56 L      B) 4.48 L      C) 1.12 L      D) 22.4 L      E) 2.24 L

ANS: C

1. Sodium hydride reacts with excess water to produce sodium hydroxide (aq) and hydrogen gas:  $NaH(s) + H_2O(l) \rightarrow NaOH(aq) + H_2(g)$

A sample of  $NaH$  weighing \_\_\_\_\_ g will produce 982 mL of gas at 28.0 °C and 765 torr, when the hydrogen is collected over water. The vapor pressure of water at this temperature is 28 torr.

- A) 2.93      B) 0.960      C) 0.925      D) 0.0388      E) 925

ANS: C

1. At STP, how many liters of oxygen are required to react completely with 3.6 liters of hydrogen to form water?  $2H_2(g) + O_2(g) \rightarrow 2H_2O(g)$

- A) 1.8 L      B) 3.6 L      C) 2.0 L      D) 2.4 L

ANS: A

1. The molar volume of a gas at STP is \_\_\_\_\_ L.

- A) 0.08206    B) 62.36    C) 1.00    D) 22.4    E) 14.7

ANS: D

1. If 0.5 L of  $O_2(g)$  reacts with  $H_2$  to produce 1 L of  $H_2O(g)$ , what is the volume of  $H_2O(g)$  obtained from 1 L of  $O_{2(g)}$ ?

- A) 0.5 L      B) 1.5 L      C) 2 L      D) 2.5 L

ANS: C

# Multiple Choice Questions On Chemical Stoichiometry

**Rachel S Tattersall**



## **Multiple Choice Questions On Chemical Stoichiometry:**

## Whispering the Techniques of Language: An Emotional Journey through **Multiple Choice Questions On Chemical Stoichiometry**

In a digitally-driven earth wherever monitors reign great and instant interaction drowns out the subtleties of language, the profound strategies and psychological nuances concealed within phrases frequently get unheard. However, situated within the pages of **Multiple Choice Questions On Chemical Stoichiometry** a captivating literary value sporting with natural thoughts, lies an exceptional journey waiting to be undertaken. Penned by an experienced wordsmith, this wonderful opus invites readers on an introspective journey, delicately unraveling the veiled truths and profound influence resonating within the cloth of every word. Within the mental depths with this touching review, we shall embark upon a sincere exploration of the book's key subjects, dissect their charming writing design, and succumb to the effective resonance it evokes serious within the recesses of readers' hearts.

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