

## BUILD Vocabulary

**ionic bond** chemical bond formed when one or more electrons are transferred from one atom to another

**ion** atom that has a positive or negative charge

**covalent bond** type of bond between atoms in which the electrons are shared

**molecule** smallest unit of most compounds that displays all the properties of that compound

**van der Waals forces** weak attraction that develops between oppositely charged regions of nearby molecules

**Preface** The prefix *co-* and *co-* mean "together." A compound is made up of elements joined together, and a covalent bond connects atoms that share electrons together. **How are covalent bonds different from ionic bonds?**

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## Chemical Bonds

**KEY QUESTION** What are the main types of chemical bonds?

The atoms in compounds are held together by chemical bonds. Bonds are formed by the electrons in the outer shell, which are called valence electrons. The main types of chemical bonds are ionic bonds and covalent bonds.

**Ionic Bonds** **Ionic bonds** are formed when one or more electrons are transferred from one atom to another. This transfer causes the atom losing the electrons to become positively charged. The atom gaining the electrons becomes negatively charged. The positively and negatively charged atoms become **ions**. The oppositely charged ions are attracted to each other and form an ionic bond.

**Covalent Bonds** Sometimes atoms share electrons rather than transferring them. These shared electrons move around the nuclei of both atoms. The bond formed by sharing electrons is called a **covalent bond**. Atoms can share between two and six electrons. The structure that results is known as a **molecule**. A molecule is the smallest unit of most compounds.

**Weak Interactions** The strongest chemical bonds are ionic or covalent. Weak interactions between atoms are also important. Within living organisms, many molecules interact only briefly. **Van der Waals forces** are one type of weak interaction between molecules. They occur when molecules are very close together.

Hydrogen bonds are another type of weak interaction. They occur between a hydrogen atom of one molecule and an oxygen or nitrogen atom of a neighboring molecule.

### Visual Reading Skill Covalent Bonds

1. What kind of chemical bond holds the water molecule together? \_\_\_\_\_
2. What particles are shared between the atoms in the molecule? \_\_\_\_\_
3. What is the other main type of chemical bond? \_\_\_\_\_
4. What is the difference between covalent and ionic bonds? \_\_\_\_\_



Water molecule ( $H_2O$ )

## Properties of Water

**READING TOOL** Cause and Effect As you read the part of the lesson under "The Water Molecule," use the table below to list the causes and effects of the properties of water. Many of the causes may be the same, but each characteristic will have a different effect. Part of the table is completed for you.

Characteristic of Water	Cause	Effect
Hydrogen Bonding		The slightly negative end of a water molecule is attracted to the slightly positive end of another water molecule.
Expands Upon Freezing		
Dissolves Many Substances	Water forms hydrogen bonds.	
Cohesion		
Adhesion		
Heat Capacity		

# Miller And Levine Biology Workbook Answers Chapter 8

**Kenneth Raymond Miller, Prentice Hall  
(School Division)**



## **Miller And Levine Biology Workbook Answers Chapter 8:**

## **Miller And Levine Biology Workbook Answers Chapter 8** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the ability of words has been evident than ever. They have the capability to inspire, provoke, and ignite change. Such is the essence of the book **Miller And Levine Biology Workbook Answers Chapter 8**, a literary masterpiece that delves deep into the significance of words and their impact on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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