

3.50 To solve this type of problem, read the problem carefully, word for word. Then try to interpret parts of the problem. For example:

- 1) Formaldehyde is an aldehyde,  $\text{H}-\overset{\text{O}}{\underset{\text{H}}{\text{C}}}-\text{H}$
- 2) Trioxane - that is, 3 formaldehydes come together to form a compound  $\text{C}_3\text{H}_6\text{O}_3$ . Because no atoms are eliminated, all of the original atoms are still present.
- 3) There are no carbonyl groups. This means that trioxane cannot contain any  $\text{C}=\text{O}$  functional groups. If you look back to Table 3.1, you can see that the only oxygen-containing functional groups that can be present are either ethers or alcohols.
- 4) A monosubstituted derivative is a compound in which one of the  $\text{H}$ 's has been replaced by a  $\text{Br}$ . Because only one monosubstituted derivative is possible, we know that there can only be one type of hydrogen in trioxane. The only possibility for trioxane is:



3.51 The highest energy conformation of hexamethane has a strain energy of  $15 \text{ kJ/mol}$ . Because this includes two  $\text{H}-\text{H}$  eclipsing interactions of  $4.0 \text{ kJ/mol}$  each, the value of an  $\text{H}-\text{H}$  eclipsing interaction is  $15 \text{ kJ/mol} - 2(4.0 \text{ kJ/mol}) = 7.0 \text{ kJ/mol}$ .

Energy

$15 \text{ kJ/mol}$

$60^\circ$



# **Mcmurry Organic Chemistry 8e Solutions Manual**

**Susan McMurry**



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